

Keywords: *work, work-life balance, work-family balance, work-family conflict, work-family interface, work-family issues, work-family research, work-family studies, work-family topics, work-family variables, work-family research topics, work-family research areas, work-family research fields, work-family research domains, work-family research subjects, work-family research issues, work-family research problems, work-family research questions, work-family research topics, work-family research areas, work-family research fields, work-family research domains, work-family research subjects, work-family research issues, work-family research problems, work-family research questions*

NUMBER 11 **ISSN 0001-9061**

Source: <http://www.fishbase.org>.

10. A sample of gas at 2.0°C and a pressure of 1.00 atm is heated to 20.0°C . What volume will this gas occupy at 1.00 atm and 20.0°C ?
11. A sample of hydrogen gas (H_2) has a volume of 10.0 L at a temperature of 10°C and a pressure of 1.00 atm . Calculate the volume of H_2 molecules present inside this sample.
12. A sample containing 100 mol of compressed gas at 1.00 atm pressure. Before this gas is released into a large plastic bag, the bag inflates to a volume of 2.00 L . What is the pressure of this gas inside the plastic bag? Assume the temperature is constant.
13. A 1.00 L gas sample at 10°C and a pressure of 100.7 kPa expands to a volume of 2.00 L . What pressure is the gas at 100.7 kPa ? What is the final temperature of the gas in degrees Celsius?
14. What is the volume of 1.00 mol of H_2 at 10°C ?
15. A hydrogen-filled balloon is a volume of 1.00 m^3 and is at a temperature of 20.0°C . The hydrogen is released at constant pressure at a temperature of 1.00 atm . What is the final volume of the balloon?
16. A 10.00 mol of nitrogen gas occupies a volume of 10.0 L at 10°C . What volume will 2.00 mol of nitrogen gas occupy at the same temperature and pressure?
17. A 1.00 L sample is filled with 1.00 g of gas. What is the composition of the gas in the 1.00 L case?
18. A particular balloon is designed to be used for the purpose of being inflated to a volume of no more than 1.00 L . If the balloon is inflated with 1.00 L of hydrogen gas at 1.00 atm , the volume of the gas is released into a large plastic bag. What is the final volume of the gas in the bag? Assume the temperature is constant.
19. A compressed gas cylinder at 100.7 kPa and 10°C is a gas cylinder that is used to fill a gas cylinder at 100.7 kPa . What is the final pressure in the cylinder?
20. A cylinder at 100.7 kPa is a volume of 1.00 L . A 1.00 g of H_2 gas and 1.00 g of O_2 gas. Calculate the total pressure of 1.00 kPa for the cylinder.
21. What volume is occupied by 1.00 mol of hydrogen gas at 10°C ?
22. Compare the diffusion rates of H_2 and O_2 gases.
23. Three moles of hydrogen gas at 1.00 L hydrogen gas at 100.7 kPa and 10°C .

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Dr. Sunita Bhagiya, ,Er. Meera Goyal

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