

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

**CHAPTER 9** STUDY GUIDE FOR CONTENT MASTERY

## Covalent Bonding

### Section 9.1 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	$\pi$ bond
---------------	----------	------------	------------	------------

When sharing of electrons occurs, the attachment between atoms that results is called a(n) **(1) covalent bond**. When such an attachment is formed, bond dissociation energy is released, and the process is **(2) exothermic**. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3) molecule**. If the electrons shared are contained between the two atoms, the attachment is called a(n) **(4) sigma bond**. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5)  $\pi$  bond**.

In your textbook, read about single and multiple bonds and bond strengths.

Circle the letter of the choice that best completes the statement or answers the question.

- Is what form do elements such as hydrogen, nitrogen, and oxygen normally occur?
  - a single atom
  - a molecule containing three atoms**
  - a molecule containing two atoms
  - a molecule containing four atoms
- How many electrons are shared in a double covalent bond?
  - none
  - one
  - two
  - four**
- Bond length is the distance between
  - two molecules of the same substance
  - the nuclei of two attached atoms**
  - the electrons in two attached atoms
  - the orbitals of two attached atoms
- Which of the following relationships relating to bond length is generally correct?
  - As the character of the bond, the stronger the bond.**
  - As the character of the bond, the weaker the bond.
  - As the character of the bond, the fewer the electrons in it.
  - As the character of the bond, the lower the bond dissociation energy.

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Chemistry: Matter and Change • Chapter 9

49

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**CHAPTER 9** STUDY GUIDE FOR CONTENT MASTERY

### Section 9.2 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write *true* or *false*.

- false** Binary molecular compounds are generally composed of a metal and a nonmetal.
- true** The second element in the formula of a binary compound is named using the suffix *-ite*.
- false** The prefix *mono-* indicates three atoms.
- true** The prefix *kilo-* indicates six atoms.
- true** In naming the first element in a formula, the prefix *mono-* is not used.
- true** For binary acids, the hydrogen part of the compound is named using the prefix *hydro-*.
- false** An oxyacid contains only two elements.
- false** If the name of the cation of an oxyacid ends in *-ite*, the acid name contains the suffix *-ate*.

In your textbook, read about naming molecular compounds and oxyacids.

For each item in Column A, write the letter of the matching item in Column B.

Column A	Column B
5. $\text{Cl}_2$	a. hydrobromic acid
6. $\text{CO}$	b. nitrogen trioxide
7. $\text{H}_2\text{O}$	c. carbon monoxide
8. $\text{H}_2\text{CO}_3$	d. nitrous acid
9. $\text{NO}_2$	e. acetone
10. $\text{N}_2\text{O}_5$	f. nitric acid
11. $\text{HNO}_3$	g. carbonic acid
12. $\text{HNO}_2$	h. bromic acid
13. $\text{HBr}$	i. carbon dioxide
14. $\text{HBrO}_3$	

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# Study Guide For Content Mastery Section 1stoichiometric Calculations

**R Barnett**



## **Study Guide For Content Mastery Section 1stoichiometric Calculations:**

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**Stoichiometric Calculations** Lifeliqe,2019 This lesson plan covers how to calculate moles and volume ratios for balanced chemical equations      *Students' Conceptions of Stoichiometry at the Submicro Level* Sin-Yan Chan,□□□,2017-01-27 This dissertation Students Conceptions of Stoichiometry at the Submicro Level by Sin yan Chan was obtained from The University of Hong Kong Pokfulam Hong Kong and is being sold pursuant to Creative Commons Attribution 3 0 Hong Kong License The content of this dissertation has not been altered in any way We have altered the formatting in order to facilitate the ease of printing and reading of the dissertation All rights not granted by the above license are retained by the author Abstract Stoichiometry is an important topic in chemistry It tells how many reactants are required to produce a certain amount of product in terms of mass mole and volume Learning stoichiometric calculation involves the understanding of certain concepts such as the mole stoichiometric ratios and chemical equations Some studies attributed the failure in learning stoichiometry to the unfamiliarity with the amount of substance in a mole and students weakness in the mathematical ability Nevertheless how students connect the submicro level and the symbolic level in learning stoichiometry was not discussed widely In this study two examples of chemical reactions with different levels of difficulties were used to probe students conceptual understanding in stoichiometry at the submicro level Their strategies used in stoichiometric calculations were examined by an interview study of five Secondary Five students The connection between the submicro level and the symbolic level in learning stoichiometry would be also probed into Results indicate that the failure of stoichiometry learning may due to the disconnection in different levels of representation and students generated strategy one portion reasoning An implication for teaching and learning is that teachers should use diagrams at the submicro level in the teaching of stoichiometry Such diagrams should aim to help students building connections across the three levels of representation and enhancing students conceptual understanding in stoichiometry DOI 10 5353 th\_b5396387 Subjects Chemistry Study and teaching Secondary

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## **Table of Contents Study Guide For Content Mastery Section 1stoichiometric Calculations**

1. Understanding the eBook Study Guide For Content Mastery Section 1stoichiometric Calculations
  - The Rise of Digital Reading Study Guide For Content Mastery Section 1stoichiometric Calculations
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Content Mastery Section 1stoichiometric Calculations
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide For Content Mastery Section 1stoichiometric Calculations
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Content Mastery Section 1stoichiometric Calculations
  - Personalized Recommendations
  - Study Guide For Content Mastery Section 1stoichiometric Calculations User Reviews and Ratings
  - Study Guide For Content Mastery Section 1stoichiometric Calculations and Bestseller Lists
5. Accessing Study Guide For Content Mastery Section 1stoichiometric Calculations Free and Paid eBooks
  - Study Guide For Content Mastery Section 1stoichiometric Calculations Public Domain eBooks
  - Study Guide For Content Mastery Section 1stoichiometric Calculations eBook Subscription Services
  - Study Guide For Content Mastery Section 1stoichiometric Calculations Budget-Friendly Options
6. Navigating Study Guide For Content Mastery Section 1stoichiometric Calculations eBook Formats
  - ePub, PDF, MOBI, and More
  - Study Guide For Content Mastery Section 1stoichiometric Calculations Compatibility with Devices
  - Study Guide For Content Mastery Section 1stoichiometric Calculations Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Study Guide For Content Mastery Section 1stoichiometric Calculations
  - Highlighting and Note-Taking Study Guide For Content Mastery Section 1stoichiometric Calculations
  - Interactive Elements Study Guide For Content Mastery Section 1stoichiometric Calculations

8. Staying Engaged with Study Guide For Content Mastery Section 1stoichiometric Calculations
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guide For Content Mastery Section 1stoichiometric Calculations
9. Balancing eBooks and Physical Books Study Guide For Content Mastery Section 1stoichiometric Calculations
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide For Content Mastery Section 1stoichiometric Calculations
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Study Guide For Content Mastery Section 1stoichiometric Calculations
  - Setting Reading Goals Study Guide For Content Mastery Section 1stoichiometric Calculations
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide For Content Mastery Section 1stoichiometric Calculations
  - Fact-Checking eBook Content of Study Guide For Content Mastery Section 1stoichiometric Calculations
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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