

# Inorganic Chemistry

What is inorganic chemistry? All of the chemistry not covered in carbon-centered organic and biochemistry—the remaining 90 or so elements!

## PERIODIC TABLE OF THE ELEMENTS

Periodic Table

### 1. Elemental Forms

- Monatomic** - Noble Gases: helium, neon, argon, krypton, xenon, radon; non-reactive
- Diatomic molecules**  
Gas: fluorine, chlorine, hydrogen, oxygen, nitrogen  
Liquid: bromine      Solid: iodine
- Molecular solid**: sulfur
- Covalent solid**: carbon (diamond), graphite
- Metals**: most common type of element  
- High coordination numbers; very dense  
- Delocalized electrons form energy bands

### 2. Types of Compounds: large variety due to the "ionic-covalent-metallic" range

- Ionic**: electron exchange
- Covalent**: shared electrons
- Polar covalent**: unequal sharing
- Ion complex**: central ion + solvent
- Transition metal compounds**  
- Coordination compound: non-bonding ligands  
- Organometallic: metal-C bond
- Complex inorganic**: metal oxides and silicates

### 3. Periodicity: chemical properties of an element depend on valence electrons Family of elements: columns in Periodic Table

#### Periodic Trends

- Chemistry determined by **valence electrons**
- Atoms or ions get a bit larger as you go down the column; electronic shells are added
- Cation and anion **charge** decrease

## REACTIONS

### Inorganic = many types of reactions

- Acid-base neutralization**
- Dissolution of electrolytes in water**
- Oxidation**: loss of electrons, add oxygen
- Reduction**: gain of electrons, remove oxygen

**TIP: Remember the acronym, OILRIG**

- Precipitation**: dissolved ions form solids
- Hydrolysis**: water reacts with a compound
- Decomposition** of compounds to elements

## CHEMISTRY OF THE METALS

### 1. Group 1: Alkali Metals

- Configuration**: noble gas + 1 electron, only valence: +1
- Natural occurrence**: ocean salts, oxide minerals
- Common magnetic**: strong base (NaOH, KOH); ionic salts (NaCl, KCl); soluble carbonate, nitrate, sulfate
- Reactions**: pure metals react violently with water; never found in nature as pure metal  
 $2Na + 2H_2O \rightarrow 2NaOH(aq) + H_2(g)$
- Spontaneous electrolysis** of molten salts  
 $NaCl(l) \rightarrow Na(l) + Cl_2(g)$
- Solubility**: water forms cage  $M \cdot H_2O$   
**Hydration**: 6 (Li<sup>+</sup>) to 8 (Cs<sup>+</sup>)  
**Hydration energy**: varies from low and weak

Li  
Na  
K  
Rb  
Cs

### 2. Group 2: Alkaline Earth Metals

- Configuration**: filled s-subshell, only valence: +2
- Natural occurrence**: mineral carbonates, silicates, ocean salts
- Reactions**:  
 $Ca(s) + 2H_2O \rightarrow Ca(OH)_2(aq) + H_2(g)$

Be  
Mg  
Ca  
Sr  
Ba

### 3. 1st Transition Series: Ti-Cu

Ti V Cr Mn Fe Co Ni Cu

- Configuration**: partially filled d-subshell; diverse chemical and electronic properties
- Periodic Trends**: electronegativity fairly constant; electrons obey Hund's Rule; density increases with Z
- Ligand Field Effect**: split d-orbitals stabilize

**APP**: white paint pigment



**APP**: attach organic groups to inorganic solids

**EX: Manganese Chemistry**

Valence ranges from 0 to 7, d<sup>5</sup> to d<sup>0</sup>

**Potassiumate**:  $MnO_4^-$ , oxidizing agent

$Mn(VII)$  easily reduced

### 4. 2nd & 3rd Transition Series: Zr-Au

|    |    |    |    |    |    |    |    |
|----|----|----|----|----|----|----|----|
| Zr | Nb | Mo | Tc | Ru | Rh | Pd | Ag |
| Hf | Ta | W  | Re | Os | Ir | Pt | Au |

- Configuration**: comparable to the 1st transition series
- Trends in radii**: "lanthanide contraction"  
- minimal change in radii for large Z; weak screening from f-orbitals
- M-M multiple bonds**: 2nd and 3rd series cluster compounds; delta (δ) bond, face-face d-orbitals

delta bond



**EX: Tetraoctate** of Mo, Rh & Ru

**d. Complex molecular structure**:  $NbF_6^+$  exists as a tetramer

**#47 Silver**: normally a non-reactive metal  
Silver reacts with S  $\rightarrow$  Ag<sub>2</sub>S - silver tarnish  
 $Ag_2O + heat \rightarrow Ag$

**#42 Molybdenum**

Quadruple M-M bond

Molybdenum  $Mo_2Cl_8^{4-}$

tetrahedral - no lone pairs

**Hallide**:  $MoF_6$ , dimer -  $Mo_2F_{10}$



**#43 Technetium**: synthetic element

### 5. Platinum Metals: Ru, Os, Rh, Ir, Pd, Pt

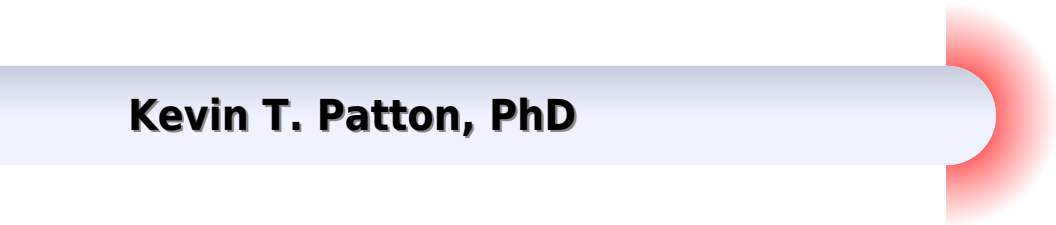
- Oxidation state** (0-6); oxygenation platinum  
**APP**: catalysts for hydrogenation; catalytic converter

### 6. Scandium, Yttrium, Lanthanum & the Lanthanides

- Lanthanide**: element with occupied f-subshell

# Study Guide For Inorganic Chemistry

**Kevin T. Patton, PhD**



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