

## Honors Chemistry

### Study Guide: Unit 5

This study guide is an overview of Unit 1 material. Please answer all of the questions to prepare for the unit test. We will work any problems you have questions about before the unit test.

All notes, BR, classwork, homework, and labs are fair game questions for the unit test.

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#### Notes

- Stoichiometry
- Mole Conversions
- Percent Composition
- Empirical vs Molecular Formula
- Heat Transfer
- Gas Laws
- Study Unit 5 Vocabulary

#### Stoichiometry

1. Define Stoichiometry
2. What is a mole?
3. What is Avogadro's Number?
4. Draw the Mole Highway
5. 5.59 g iron = \_\_\_\_\_ moles iron
6. 2.5 moles of hydrogen = \_\_\_\_\_ grams hydrogen
7. 0.100 moles of a compound has a mass of 6.539 g. Formula mass = \_\_\_\_\_

#### Percent Composition

8. What is percent composition?
9. Write the equation for percent composition.
10. List the steps for finding percent composition.
11. What is the percent composition of hydrogen and oxygen in hydrogen peroxide,  $\text{H}_2\text{O}_2$ ?

#### Empirical vs Molecular Formula

12. Study Cheat Sheet
13. What is the difference between Empirical and Molecular formulas?
14. List the steps to find Empirical Formulas from percent's of each element.
15. List the steps to find Molecular Formulas from an Empirical Formula.

#### Heat Transfer

16. The amount of energy required to raise the temp of 1 kg of material by 1 degree Kelvin is called?
17. The amount of energy required to convert solid at its melting point into a liquid without an increase in temperature is called?
18. The amount of heat required to convert a liquid into a gas without increasing the temperature is called?
19. In a heating and cooling diagram, what does a plateau show?
20. What is endothermic?
21. What is exothermic?
22. A device used to measure changes in thermal energy is called?

# Study Guide Honors Chemistry Answer

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