

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

# CHAPTER 3 STUDY GUIDE FOR CONTENT MASTERY

## Matter—Properties and Changes

### Section 3.1 Properties of Matter

In your textbook, read about physical properties and chemical properties of matter.

Use each of the terms below just once to complete the passage.

|                     |                    |                       |
|---------------------|--------------------|-----------------------|
| chemical<br>density | mass<br>properties | physical<br>substance |
|---------------------|--------------------|-----------------------|

Matter is anything with (1) mass and volume. A (2) substance is a form of matter with a uniform and unchanging composition. Substances have specific, unchanging (3) properties that can be observed. Substances have both physical and chemical properties. (4) Physical properties can be observed without changing a substance's chemical composition. Color, hardness, and (5) density are examples. Other properties cannot be observed without changing the composition of a substance. These are called (6) chemical properties. An example is the tendency of iron to form rust when exposed to air.

Label each property as either physical or chemical.

- |                 |                                                           |
|-----------------|-----------------------------------------------------------|
| <u>physical</u> | 7. Chemical formula $H_2O$                                |
| <u>chemical</u> | 8. Irons react vigorously when exposed to moist air       |
| <u>chemical</u> | 9. Remains unchanged when in the presence of nitrogen     |
| <u>physical</u> | 10. Colorless                                             |
| <u>physical</u> | 11. Solid at normal temperatures and pressures            |
| <u>chemical</u> | 12. Ability to combine with another substance             |
| <u>physical</u> | 13. Melting point                                         |
| <u>physical</u> | 14. Liquid at normal temperatures and pressures           |
| <u>physical</u> | 15. Boiling point is $100^\circ\text{C}$                  |
| <u>physical</u> | 16. Conducts electricity                                  |
| <u>physical</u> | 17. Density is $\frac{1}{2} \frac{\text{g}}{\text{cm}^3}$ |

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# CHAPTER 3 STUDY GUIDE FOR CONTENT MASTERY

### Section 3.1 continued

In your textbook, read about states of matter.

Label each drawing with one of these words: solid, liquid, gas.



For each statement below, write true or false.

- |              |                                                                                          |
|--------------|------------------------------------------------------------------------------------------|
| <u>true</u>  | 21. All matter that we encounter in everyday life exists in one of these physical forms. |
| <u>true</u>  | 22. A solid has definite shape and volume.                                               |
| <u>false</u> | 23. A liquid has a definite shape and takes on the volume of its container.              |
| <u>true</u>  | 24. A gas has both the shape and the volume of its container.                            |
| <u>false</u> | 25. The particles in a gas cannot be compressed into a smaller volume.                   |
| <u>false</u> | 26. Liquids tend to contract when heated.                                                |
| <u>false</u> | 27. The particles in a solid are spaced far apart.                                       |
| <u>false</u> | 28. The words gas and vapor can be used interchangeably.                                 |

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