

Name: \_\_\_\_\_

## Atomic Structure & the Periodic Table Study Guide

### Atomic Structure

Particle	Charge	Where are they found?
Proton	Positive	Nucleus
Neutron	Neutral/ no charge	Nucleus
Electron	negative	Electron cloud/energy levels outside nucleus

1. What is the rule for writing chemical symbols like Na or Ca? **Capitalize the first letter; lower case second letter**
2. What are the vertical columns on the periodic table called? **Groups** or **families**
  - a. How many of these are on the periodic table? **18**
  - b. Why are the elements in these columns? **They have the same chemical properties** (& same # of valence electrons)
  - c. Elements in the same group have the same number of **valence** electrons.
3. What are the horizontal rows on the periodic table called? **periods**
  - a. How many of these are on the periodic table? **7**
  - b. Elements in these rows have the same number of **energy** (electron shells) levels.
4. What are valence electrons? **Electrons in the outer energy level**
5. How many electrons are in each energy level?  
1<sup>st</sup> energy level **2**    2<sup>nd</sup> energy level **8**    3<sup>rd</sup> energy level **8**
6. The atomic number is equal to the **# of protons** and also **# of electrons**
7. The atomic mass is equal to the **# protons + # of neutrons**
8. How do you find the number of neutrons? **Atomic mass – atomic number**
9. We view the order of the periodic table today according to the **atomic number**

# Study Guide Periodic Trends

**The Princeton Review**



## **Study Guide Periodic Trends:**

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement ASAP Chemistry: A Quick-Review Study Guide for the AP Exam The Princeton Review, 2019-02-12 Looking for sample exams practice questions and test taking strategies Check out our extended in depth AP chem prep guide Cracking the AP Chemistry Exam LIKE CLASS NOTES ONLY BETTER The Princeton Review s ASAP Chemistry is designed to help you zero in on just the information you need to know to successfully grapple with the AP test No questions no drills just review Advanced Placement exams require students to have a firm grasp of content you can t bluff or even logic your way to a 5 Like a set of class notes borrowed from the smartest student in your grade this book gives you exactly that No tricks or crazy stratagems no sample essays or practice sets Just the facts presented with lots of helpful visuals Inside ASAP Chemistry you ll find Essential concepts terms and functions for AP Chem all explained clearly concisely Diagrams charts and graphs for quick visual reference A three pass icon system designed to help you prioritize learning what you MUST SHOULD and COULD know in the time you have available Ask Yourself questions to help identify areas where you might need extra attention A resource that s perfect for last minute exam prep and for daily class work Topics covered in ASAP Chemistry include Atomic structure Covalent bonding intermolecular forces Thermochemistry Acids bases and more

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