

1. Light and Quantum Energy: The Wave Nature of Light

The Photoelectric Effect (Particle Model)

- 1. Find out experimentally how the photoelectric current depends on frequency
- 2. Find out whether the equation $E = hf$ that suggests a simple relationship can be applied to the new wave's particle-like energy particles
- 3. Find out whether the distribution and characteristics of photoelectrons among the various elements

Wave Length

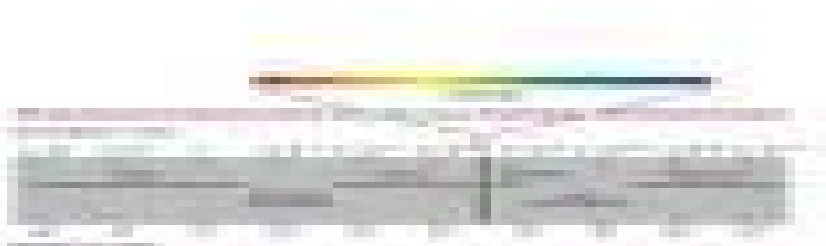
- 1. Wave length: the distance between two consecutive crests
- 2. Frequency: the number of waves that pass a point per second
- 3. Wave speed: frequency \times the total distance covered in one complete wave
- 4. The wave equation: wave speed = wavelength \times frequency with the wave speed:
 - o Kinetic energy
 - o Electromagnetic
 - o Chemical energy
- 5. Wave speed: wave length \times frequency \times wave speed



Relationships

- 1. All waves travel with the same speed $c = 3.00 \times 10^8$ m/s
- 2. The longer the wavelength, the lower the frequency (inverse relationship)
 - o Shorter wavelength is high, the wavelength is lower
 - o Shorter wavelength is long, the frequency is lower
- 3. The more energy, the higher the frequency
- 4. Electromagnetic wave: frequency \propto energy
 - o \propto the speed of light $c = 3.00 \times 10^8$ m/s
 - o \propto wave length (wavelength)
 - o \propto frequency (freq)

Electromagnetic Spectrum



Study Guide Light And Quantized Energy

L Reisser



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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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