

Chemistry I Term 1 exam sample problems by topic

physical or chemical change

- boiling water P
- burning wood P
- melting aluminum P
- separating sand from water P

element names/symbols - know element symbols from handout during 1st week of class

metric prefixes (centi-, milli-, kilo-)

- Convert 54 mm to m.
- Convert 9.2×10^5 cL to L.
- Convert 5.12 kg to g.

$$\frac{54 \cancel{\text{mm}}}{1000 \cancel{\text{mm}}} = 0.054 \text{ m}$$

$$\frac{9.2 \times 10^5 \cancel{\text{cL}}}{100 \cancel{\text{cL}}} = 9.2 \times 10^3 \text{ L}$$

$$\frac{5.12 \text{ kg}}{1000 \text{ g}} = 5120 \text{ g}$$

Temperature conversions

- Convert 23 degrees C to Kelvin
- Convert 50 K to degrees C
- Convert 37 degrees C to degrees F

$$23 + 273 = 296 \text{ K}$$

$$50 - 273 = -223 \text{ C}$$

$$\frac{9}{5}(37 \text{ C}) + 32 = 98.6 \text{ F}$$

significant digits

- Round 8.9456 to 3 significant digits.
- Round 513431 to 2 significant digits.
- Add $4.324 + 345.1$, and round your answer properly.
- Multiply 32.56×103 , and round your answer properly.

$$8.95$$

$$510000$$

$$349.424 \rightarrow 349.4 \text{ (1 decimal place)}$$

$$3353.68 \rightarrow 3350 \text{ (3 sig figs)}$$

scientific notation

- Express 3712 in scientific notation.
- Express 0.0377561 in scientific notation.

$$3.712 \times 10^3$$

$$3.77561 \times 10^{-2}$$

density

What is the volume of a solid with a density of 7.8 g/mL and a mass of 11.8 g?

$$D = \frac{m}{V} \quad 7.8 \text{ g/mL} = \frac{11.8 \text{ g}}{V}$$

$$V = 1.51 \text{ mL}$$

unit conversions (inches to cm, seconds to hours, etc) including imperial and metric

- Convert 54 hours to seconds
- Convert 84 oz to grams
- Convert 1234 m to inches
- Convert 2.3 quarts to Liters

$$\frac{54 \text{ hr}}{1 \text{ hr}} \times \frac{60 \text{ min}}{1 \text{ hr}} \times \frac{60 \text{ sec}}{1 \text{ min}} = 194400 \text{ sec}$$

$$\frac{84 \text{ oz}}{16 \text{ oz}} \times \frac{454 \text{ g}}{1 \text{ lb}} = 2400 \text{ g}$$

$$\frac{2.3 \text{ qt}}{1.06 \text{ qt}} = 2.16 \text{ L}$$

$$\frac{1234 \text{ m}}{1 \text{ m}} \times \frac{100 \text{ cm}}{1 \text{ m}} \times \frac{1 \text{ inch}}{2.54 \text{ cm}} = 48600 \text{ inches}$$

Counting protons, neutrons, and electrons

Symbol	Protons	Neutrons	Electrons	Charge
${}^9_4\text{Be}$	4	5	4	0
${}^{198}_{79}\text{Au}$	47	61	47	0
${}^{201}_{80}\text{Hg}$	80	121	80	0
${}^{33}_{16}\text{S}^{2-}$	16	17	18	-2
${}^{13}_{13}\text{Al}^{+3}$	13	15	10	+3
${}^{36}_{34}\text{Se}^{-}$	34	47	35	-1
${}^{58}_{28}\text{Ni}$	28	32	28	0
${}^{80}_{34}\text{Se}^{2-}$	34	46	36	-2
${}^{16}_{8}\text{O}^{2-}$	8	8	10	-2

Study Guide Mixed Review Answer Key

Brielle Morgan



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chapter reinforce key concepts and guide you in application to evidence based clinical practice NEW Increased emphasis on evidence based practice corresponding to the EBP emphasis in the text to help you see the value of understanding the research process and apply it to evidence based nursing practice NEW Hands on practice with three current high quality published studies to help you better understand the research and evidence based practice processes and help you learn to appraise and apply research findings to clinical settings NEW Enhanced key terms activities compensate for the deletion of the key terms lists from the textbook that are now addressed in Study Guide activities NEW New Appraisal Guidelines help you to critically appraise research articles NEW Updated full text articles ensure that the examples provided reflect the most current high quality studies that are meaningful

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