



Biology

Physics

Notes



Energy is the ability to do work or transfer heat. It is a scalar quantity.



The **first law of thermodynamics** states that energy is conserved. It cannot be created or destroyed, only transformed.



For a system, the **change in internal energy** is equal to the heat added to the system minus the work done by the system.



A **closed system** is one in which no matter or energy can enter or leave. The **first law of thermodynamics** applies to closed systems.

Energy



The **second law of thermodynamics** states that the total entropy of a closed system always increases over time.

- (a) Entropy is a measure of disorder.
- (b) Entropy is a state function.
- (c) Entropy is always increasing.



Energy is the ability to do work or transfer heat. It is a scalar quantity.



The **first law of thermodynamics** states that energy is conserved. It cannot be created or destroyed, only transformed.



The **second law of thermodynamics** states that the total entropy of a closed system always increases over time.



The **third law of thermodynamics** states that the entropy of a perfect crystal at absolute zero is zero.

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- (c) Entropy is always increasing.



The **fourth law of thermodynamics** states that the temperature of a system is proportional to the average kinetic energy of its particles.

- (a) Temperature is a measure of the average kinetic energy of particles.
- (b) Temperature is a state function.
- (c) Temperature is always increasing.



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