

Chemistry 30

Unit 4: Solutions

Assignment 3 Calculations Involving Solution and Ion Concentrations

You must clearly show all work for your calculations. Be sure to highlight your final answer – circle it, underline it, whatever. Be sure to include the units for your answer. Failure to follow these directions will result in a loss of marks.

1. What mass of sodium hydroxide is needed to make 500.0 mL of 0.5 M NaOH solution.

$$\text{Molar Mass of NaOH} = 40.0 \text{ g}\cdot\text{mol}^{-1}$$

$$g = \frac{40.0g}{\text{mol}} \times \frac{0.5\text{mol}}{L} \times \frac{0.500L}{1} = 10.0g$$

2. What volume of 0.060 M KCl solution contains 2.39 g of KCl?

$$\text{Molar mass of KCl} = 74.6 \text{ g}\cdot\text{mol}^{-1}$$

$$L = \frac{L}{0.060\text{mol}} \times \frac{\text{mol}}{74.6g} \times \frac{2.39g}{1} = 0.53L \text{ or } 530 \text{ mL}$$

3. Calculate the concentration of a solution prepared by dissolving 24.2 g of $\text{Fe}_2(\text{SO}_4)_3$ in enough water to make 250.0 mL of solution.

$$\text{Molar mass of Fe}_2(\text{SO}_4)_3 = 399.9 \text{ g}\cdot\text{mol}^{-1}$$

$$M = \frac{\text{mol}}{L} = \frac{\text{mol}}{399.9g} \times \frac{24.2g}{1} \times \frac{1}{0.250L} = 0.242M$$

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Arshad Iqbal



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Table of Contents Study Guide For Solution Concentration Answer Key

1. Understanding the eBook Study Guide For Solution Concentration Answer Key
 - The Rise of Digital Reading Study Guide For Solution Concentration Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Solution Concentration Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide For Solution Concentration Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Solution Concentration Answer Key
 - Personalized Recommendations
 - Study Guide For Solution Concentration Answer Key User Reviews and Ratings

- Study Guide For Solution Concentration Answer Key and Bestseller Lists
- 5. Accessing Study Guide For Solution Concentration Answer Key Free and Paid eBooks
 - Study Guide For Solution Concentration Answer Key Public Domain eBooks
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- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide For Solution Concentration Answer Key
 - Highlighting and Note-Taking Study Guide For Solution Concentration Answer Key
 - Interactive Elements Study Guide For Solution Concentration Answer Key
- 8. Staying Engaged with Study Guide For Solution Concentration Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide For Solution Concentration Answer Key
- 9. Balancing eBooks and Physical Books Study Guide For Solution Concentration Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide For Solution Concentration Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide For Solution Concentration Answer Key
 - Setting Reading Goals Study Guide For Solution Concentration Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide For Solution Concentration Answer Key
 - Fact-Checking eBook Content of Study Guide For Solution Concentration Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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