

Molarity Practice

Name: _____

Date: _____

1. Calculate the molarity of the following solutions:

a) 45 g of Na_2SO_4 in 150 mL of solution.

$$(45\text{g}) \left(\frac{1 \text{ mol}}{142.1\text{g}} \right) \left(\frac{1}{0.150\text{L}} \right) = \boxed{2.1 \text{ M}}$$

b) 24.6 g of $(\text{NH}_4)_2\text{CO}_3$ in 75 mL of solution.

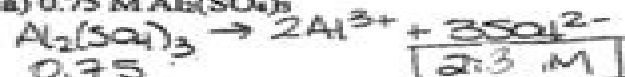
$$(24.6\text{g}) \left(\frac{1 \text{ mol}}{96.1\text{g}} \right) \left(\frac{1}{0.075\text{L}} \right) = \boxed{3.4 \text{ M}}$$

c) 73.1 g of $\text{Ca}(\text{NO}_3)_2$ in 125 mL of solution.

$$(73.1\text{g}) \left(\frac{1 \text{ mol}}{164.1\text{g}} \right) \left(\frac{1}{0.125\text{L}} \right) = \boxed{3.6 \text{ M}}$$

2. What is the concentration of sulphate ions, SO_4^{2-} , in each of the following?

a) 0.75 M $\text{Al}_2(\text{SO}_4)_3$



b) 1.35 M Na_2SO_4



3. What is the molarity of chlorine ions in solution when 47 g of AlCl_3 is dissolved in a 210 mL of solution?

$$(47\text{g}) \left(\frac{1 \text{ mol}}{133.5\text{g}} \right) \left(\frac{1}{0.210\text{L}} \right) = 1.68 \text{ M}$$

$$[\text{Cl}^-] = 3(1.68\text{M}) = \boxed{5.0 \text{ M}}$$

4. Which of the following solutions has the highest concentration? Prove using calculations.

A) 12.5 g of CaCl_2 in 40 mL of solution

$$(12.5\text{g}) \left(\frac{1 \text{ mol}}{111.1\text{g}} \right) \left(\frac{1}{0.040\text{L}} \right) = 2.82 \text{ M}$$

B) 20.9 g of MgI_2 in 35 mL

$$(20.9\text{g}) \left(\frac{1 \text{ mol}}{278.1\text{g}} \right) \left(\frac{1}{0.035\text{L}} \right) = 2.2 \text{ M}$$

5. How many grams of salt (NaCl) need to be dissolved in 300 mL of solution to give you a solution that has a concentration of 1.2 M? (Hint: you need to work backwards on this one. You are given the molarity and the liters, so find the moles and convert to grams.)

$$\left(\frac{1.2 \text{ mol}}{\text{L}} \right) (0.300\text{L}) = 0.36 \text{ mol}$$

$$\text{m} = (0.36 \text{ mol}) \left(\frac{58.5\text{g}}{1} \right) = \boxed{21. \text{ g}}$$

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