

Electron Configuration Practice Worksheet **KEY**

In the space below, write the unabbreviated electron configurations of the following elements:

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^4$

In the space below, write the abbreviated electron configurations of the following elements:

- 6) cobalt $[Ar] 4s^2 3d^7$
- 7) silver $[Kr] 5s^2 4d^9$
- 8) tellurium $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium $[Rn] 7s^2$
- 10) lawrencium $[Rn] 7s^2 5f^{14} 6d^1$

Determine what elements are denoted by the following electron configurations:

- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ **Sulfur**
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ **Rubidium**
- 13) $[Kr] 5s^2 4d^{10} 5p^3$ **Antimony**
- 14) $[Xe] 6s^2 4f^{14} 5d^8$ **Osmium**
- 15) $[Rn] 7s^2 5f^{11}$ **Einsteinium**

Determine which of the following electron configurations are not valid:

- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ **should be $3d^{10}$**
- 17) $1s^2 2s^2 2p^6 3s^2 3d^5$ **$3s^2$ is impossible & "p" instead of "d"**
- 18) $[Ra] 7s^2 5f^8$ **Ra is not a noble gas**
- 19) $[Kr] 5s^2 4d^{10} 5p^5$ **Valid**
- 20) $[Xe]$ **Element can't be its own electron configuration.**

Unit 4 Electrons Answer Key

Mr. Rohit Manglik



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Table of Contents Unit 4 Electrons Answer Key

1. Understanding the eBook Unit 4 Electrons Answer Key
 - The Rise of Digital Reading Unit 4 Electrons Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 4 Electrons Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 4 Electrons Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 4 Electrons Answer Key
 - Personalized Recommendations

- Unit 4 Electrons Answer Key User Reviews and Ratings
- Unit 4 Electrons Answer Key and Bestseller Lists
- 5. Accessing Unit 4 Electrons Answer Key Free and Paid eBooks
 - Unit 4 Electrons Answer Key Public Domain eBooks
 - Unit 4 Electrons Answer Key eBook Subscription Services
 - Unit 4 Electrons Answer Key Budget-Friendly Options
- 6. Navigating Unit 4 Electrons Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 4 Electrons Answer Key Compatibility with Devices
 - Unit 4 Electrons Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 4 Electrons Answer Key
 - Highlighting and Note-Taking Unit 4 Electrons Answer Key
 - Interactive Elements Unit 4 Electrons Answer Key
- 8. Staying Engaged with Unit 4 Electrons Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 4 Electrons Answer Key
- 9. Balancing eBooks and Physical Books Unit 4 Electrons Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 4 Electrons Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 4 Electrons Answer Key
 - Setting Reading Goals Unit 4 Electrons Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 4 Electrons Answer Key
 - Fact-Checking eBook Content of Unit 4 Electrons Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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