

Unit 9 Worksheet: St. Michael's Concentration

Part 9: Describing Molar Concentrations

1. Sodium chloride was dissolved in water to produce a 1.5M solution. Suppose you had three separate beakers, each about the size of a glass, that weighed as indicated. How much sodium chloride would you add to each?
 These questions sometimes (1.5) involve **NaCl** for answer (1.5) and sometimes, if needed, **NaCl** and the amount **NaCl** would be added (1.5) to the solution.
2. A 10.0 g solution of potassium chloride is dissolved in enough water to make 100.0 mL of solution. Calculate the mass percent of potassium chloride.

$$10.0 \text{ g} = \frac{1.00 \text{ mol}}{1.00 \text{ mol}} \times 1.00 \text{ mol} = 1.00 \text{ mol}$$

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3. How much sodium chloride is dissolved in enough water to make 100.0 g of solution? Calculate the mass percent of sodium chloride.

$$10.0 \text{ g} = \frac{1.00 \text{ mol}}{1.00 \text{ mol}} \times 1.00 \text{ mol} = 1.00 \text{ mol}$$

$$10.0 \text{ g} = \frac{1.00 \text{ mol}}{1.00 \text{ mol}} \times 1.00 \text{ mol} = 1.00 \text{ mol}$$

4. A solution of sodium chloride is prepared by dissolving 10.0 g of sodium chloride in enough water to make 100.0 mL of solution. Calculate the mass percent of sodium chloride.

$$10.0 \text{ g} = \frac{1.00 \text{ mol}}{1.00 \text{ mol}} \times 1.00 \text{ mol} = 1.00 \text{ mol}$$

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5. A solution of sodium chloride is prepared by dissolving 10.0 g of sodium chloride in enough water to make 100.0 mL of solution. Calculate the mass percent of sodium chloride.

$$10.0 \text{ g} = \frac{1.00 \text{ mol}}{1.00 \text{ mol}} \times 1.00 \text{ mol} = 1.00 \text{ mol}$$

Unit 11 Worksheet 3 Molar Concentration

O. Jerry Parker, Gary L. Breneman

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the authors' names.

Unit 11 Worksheet 3 Molar Concentration:

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