

Stoichiometry:
Index for Stoichiometry Practice

Name _____

Questions 1 – 3 are based on the following chemical equation:



1. How many moles of H_2 can be produced from 1.8 moles Al ?
2. If 3.00 moles of H_2SO_4 reacts with excess Al , how many moles of $\text{Al}_2(\text{SO}_4)_3$ will be produced?
3. How many moles of aluminum will be needed to produce 10.0 moles of hydrogen gas?

Questions 4 – 6 are based on the decomposition of potassium chlorate. When heated, potassium chlorate decomposes into potassium chloride and oxygen.

Balance Equation: _____

4. How many moles of O_2 can be produced from the decomposition of 1.00 moles KClO_3 ?
5. If 6.000 moles of KClO_3 decomposes, how many moles of KCl will be produced?
6. How many moles of potassium chloride will be needed to produce 10.0 moles of oxygen gas?

Challenge: How many moles of oxygen gas will be produced from the decomposition of 20.0 grams of KClO_3 ?

Skill Practice 31 Moles And Formulas

Peter Atkins, Loretta Jones



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