

Name: _____ Date: _____ Mode: _____

Determining Molecular Formulas (True Formulas)

1. The empirical formula of a compound is NO_2 . Its molecular mass is 92 g/mol. What is its molecular formula?

$$\begin{array}{r} \text{NO}_2 \\ 16.00\text{g} \\ 16.00\text{g} \\ + 14.00\text{g} \\ \hline 46.00\text{g} \end{array}$$

$$\frac{92\text{g}}{46.0\text{g}} = 2 \quad \text{N}_2\text{O}_4$$

2. The empirical formula of a compound is CH_2 . Its molecular mass is 70 g/mol. What is its molecular formula?

$$\begin{array}{r} \text{CH}_2 \\ 12.01\text{g} \\ 1.01\text{g} \\ + 1.01\text{g} \\ \hline 14.03\text{g} \end{array}$$

$$\frac{70\text{g}}{14.03\text{g}} = 4.98 = 5 \quad \text{C}_5\text{H}_{10}$$

3. A compound is found to be 40.0% carbon, 6.7% hydrogen and 53.3% oxygen. Its molecular mass is 60 g/mol. What is its molecular formula?

$$40.0\text{g C} \cdot \frac{1\text{mol C}}{12.01\text{g C}} = 3.33\text{mol C} \div 3.33 = 1$$

$$\text{CH}_2\text{O} = 30.02\text{g}$$

$$6.7\text{g H} \cdot \frac{1\text{mol H}}{1.01\text{g H}} = 6.63\text{mol H} \div 3.33 = 2$$

$$\frac{60\text{g}}{30\text{g}} = 2$$

$$53.3\text{g O} \cdot \frac{1\text{mol O}}{16.00\text{g O}} = 3.34\text{mol O} \div 3.33 = 1$$

$$\text{C}_2\text{H}_4\text{O}_2$$

4. A compound is 64.9% carbon, 13.5% hydrogen and 21.6% oxygen. Its molecular mass is 74 g/mol. What is its molecular formula?

$$64.9\text{g C} \cdot \frac{1\text{mol C}}{12.01\text{g C}} = 5.40\text{mol C} \div 1.35 = 4$$

$$\text{C}_4\text{H}_{10}\text{O}$$

$$21.6\text{g O} \cdot \frac{1\text{mol O}}{16.00\text{g O}} = 1.35\text{mol O} \div 1.35 = 1$$

↑
empirical
↑
molecular

$$13.5\text{g H} \cdot \frac{1\text{mol H}}{1.01\text{g H}} = 13.4\text{mol H} \div 1.35 = 10$$

5. A compound is 54.5% carbon, 9.1% hydrogen and 36.4% oxygen. Its molecular mass is 88 g/mol. What is its molecular formula?

$$54.5\text{g C} \cdot \frac{1\text{mol C}}{12.01\text{g C}} = 4.53\text{mol C} \div 2.28 = 2$$

$$\text{C}_2\text{H}_4\text{O} = 44$$

$$9.1\text{g H} \cdot \frac{1\text{mol H}}{1.01\text{g H}} = 9.00\text{mol H} \div 2.28 = 4$$

$$\frac{88}{2} = 44$$

$$36.4\text{g O} \cdot \frac{1\text{mol O}}{16.00\text{g O}} = 2.28\text{mol O} \div 2.28 = 1$$

$$\text{C}_4\text{H}_8\text{O}_2$$

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