

Name _____

Chemistry Period _____

Worksheet Specific Heat

Specific Heat of water = $4.18 \text{ J/g}^\circ\text{C}$

1. A sample of mercury is heated from 25.5°C to 52.5°C . In the process, 3050J of heat are absorbed. What mass of mercury was in the sample? The specific heat of mercury is $0.140 \text{ J/g}^\circ\text{C}$. **ANS: 807 g**
2. A block of aluminum with a mass of 140g is cooled from 98.4°C to 62.2°C with a release of 1137J of heat. From these data, calculate the specific heat of aluminum. **ANS: 0.224**
3. A total of 226 J of heat are absorbed as 58.3 g of lead is heated from 12.0°C to 42.0°C . From this data, what is the specific heat of lead? **ANS: 0.129**
4. What is the change in heat energy when 132g of copper at 32.2°C is raised to 45.0°C ? The specific heat of copper is $0.385 \text{ J/g}^\circ\text{C}$. **ANS: 650 J**
5. What is the change in heat energy when 40.0g of water at 25.0°C is raised to 60.0°C ? **ANS: 5852 J**
6. What is the change in heat energy when 75.0 of water at 9.8°C is raised to 22.4°C ? **3950 J**
7. What is the change in heat energy when 125g of water at 10.0°C is raised to 90.0°C ? **ANS: 41.8 kJ**
8. What is the change in heat energy when 64.82 g of aluminum metal at 100.0°C is cooled to 82.0°C ? The specific heat of aluminum is $0.897 \text{ J/g}^\circ\text{C}$. **ANS: 1047**
9. What is the change in temperature if 24g of water at 50°C absorbs 985 J of heat energy? **ANS: 9.82**
10. What is the mass of a sample of iron if that sample lost 2300J of heat energy when it cooled from 80°C to 30°C ? The specific heat of iron is $0.449 \text{ J/g}^\circ\text{C}$. **ANS: 102 g**
11. What is the specific heat of metal if its mass is 26.86g and it requires 418.6 J of heat energy to raise its temperature from 27.4°C to 67.3°C ? **ANS: 0.391**
12. How much heat energy is needed to raise the temperature of 59.7g of cadmium from 25°C to 100°C ? The specific heat of cadmium is $0.231 \text{ J/g}^\circ\text{C}$. **ANS: 1034 J**

Skills Math Skills Specific Heat Answers

Robert Jansen, Greg Young



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Math Skills for the Sciences John G. Pearson, Dennis M. Stone, Richard F. Swindell, 1976 Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction **Science Spectrum** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2003-03

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tool and three full length practice tests with detailed answer explanations **Student Edition**

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