

Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

- Write the balanced chemical equation.
- Identify what is given (with units) and what you want to find (with units).
- Use coefficients from balanced equation to determine mole ratio.
- Show set up (organize it!).

- Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?

Equation: $\underline{\quad 2 \quad} \text{H}_2\text{S}_{(g)} + \underline{\quad 3 \quad} \text{O}_{2(g)} \rightarrow \underline{\quad 2 \quad} \text{SO}_{2(g)} + \underline{\quad 2 \quad} \text{H}_2\text{O}_{(g)}$

Before $\underline{\quad 2 \quad}$ $\underline{\text{ex}}$ $\underline{\quad 0 \quad}$ $\underline{\quad 0 \quad}$ 3.0 mol O₂

Change $\underline{\quad -2 \quad}$ $\underline{\quad -3 \quad}$ $\underline{\quad 2 \quad}$ $\underline{\quad 2 \quad}$

After $\underline{\quad 0 \quad}$ $\underline{\text{ex}}$ $\underline{\quad 2 \quad}$ $\underline{\quad 2 \quad}$

- Propane, C₃H₈, burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?

Equation: $2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$

Before 0.5 $\underline{\text{ex}}$ 0 0 1.50mol HCL

Change -0.5 -1.5 +0.5 +0.5

After 0 $\underline{\text{ex}}$ 0.5 0.75

- Ammonia, NH₃, for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?

Equation: $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$

Before $\underline{\text{ex}}$ 1.25 0 2.50 mole

Change -3.75 -1.25 +2.5

After $\underline{\text{ex}}$ 0 2.5

Unit 7 Mole Relationships

Morris Hein, Susan Arena, Cary Willard

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the authors' names.

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Basic Concepts of Chemistry Leo J. Malone, Theodore O. Dolter, 2011-12-27 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. WileyPLUS sold separately from text. *Cambridge International AS and A Level Chemistry Revision Guide* Judith Potter, Peter Cann, 2015-10-29 A revision guide tailored to the AS and A Level Chemistry syllabus 9701 for first examination in 2016. This Revision Guide offers support for students as they prepare for their AS and A Level Chemistry 9701 exams. Containing up-to-date material that matches the syllabus for examination from 2016 and packed full of guidance such as Worked Examples, Tips, and Progress Check questions throughout to help students to hone their revision and exam technique and avoid common mistakes. These features have been specifically designed to help students apply their knowledge in exams. Written in a clear and straightforward tone, this Revision Guide is perfect for international learners. Foundations of College Chemistry

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Fraenkel-Conrat, 2013-04-17 The time seems ripe for a critical compendium of that segment of the biological universe we call

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