

charge on the carbon makes it polar.

Electrophile is an **electron-pair donor**.
 Electrophile is an electron pair to somewhere without enough electrons.
 H^+ and NH_3 are all nucleophiles that can react with haloalkanes.



Haloalkanes Can Undergo Nucleophilic Substitution Reactions

Haloalkane can react with a polar molecule by kicking out the functional group and taking its place.
 called a **nucleophilic substitution reaction**. It works like this:



The lone pair of electrons on the nucleophile attacks the δ^+ carbon.
 The $\text{C}-\text{X}$ bond breaks.

The halogen leaves, taking both electrons with it. A new bond forms between the carbon and the nucleophile.

Product of these reactions depends on what the nucleophile is...

Haloalkanes React with Hydroxides to Form Alcohols

Haloalkanes, bromoethane can react to form ethanol. You have to use warm sodium or potassium hydroxide — it's a nucleophilic substitution reaction.

X stands for one of the halogens (F, Cl, Br or I).
 Nu stands for a nucleophile.

This is sometimes hydrolysis, but exactly the same will happen with (Hydrolysis is splitting a molecule)

Unit 3 Section 1 Quiz Chemistry

SJ Ball



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