

Name:

Period:

Stoichiometry Unit Study Guide

Vocabulary section

Directions: Use the following words from the word bank to fill in the blanks that best fits the sentence/definitions. Each word is only used **ONCE**.

Word bank			
Stoichiometry	Grams	Molar Mass	Mole ratio
Mole number	Scientific notation	Law of conservation of mass	Balanced chemical equation

1. The _____ states that in a chemical reaction, mass is **not** created or destroyed. It also means that if an element is present among the reactants it must almost be present in the product.
2. _____ is the mass of one mole of any atom/molecule and has a unit of **g/mol**.
3. _____ is the unit that measures mass only.
4. _____ is the relationship between the amounts/quantities of substances taking part in a chemical reaction. We use the "train track method" to solve these types of problems.
5. The _____ is **1 mole = 6.02×10^{23} atoms/molecules**.
6. A _____ is like the **recipe** of a chemical reaction and must follow the law of conservation.
7. _____ is a way of writing very large or very small numbers to use them more effectively.
8. _____ is a **conversion factor** that relates the amounts in moles of any two substances in a chemical reaction.

Directions: Answer the following questions using dimensional analysis (AKA the train track method). Show all your work and include units!

1. Find the number of moles in 46.32 grams of H_2O .

Unit Stoichiometry Study Guide Answers

**John Krenos, Joseph Potenza, Carl
Hoeger**



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