

8 Structure and nomenclature of organic compounds

Unit	4	How are organic compounds categorised, analysed and used?
Area of Study	1	How can the diversity of carbon compounds be explained and categorised?
Topic	8	Structure and nomenclature of organic compounds
Subtopic	8.2	The carbon atom

Summary

- Carbon atoms have 4 valence electrons and make 4 covalent bonds with other carbon atoms or atoms of different elements.
- Bond energy is the amount of energy required to break the bonds of reactants or the amount of energy released per mole when a covalent bond forms between 2 molecules. It is typically measure in kJ mol^{-1} .
- The bond length is the distance between the nuclei of the atoms sharing the electrons.
- Carbon molecules with single bonds typically form a tetrahedral arrangement.
- Double $\text{C}=\text{C}$ bonds are stronger than $\text{C}-\text{C}$ single bonds and forms flat molecules. $\text{C}\equiv\text{C}$ bonds are stronger again and form linear molecules.



My notes

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