#### Dr. - Heder Mass

- 70. Come Ser the following:
  - a. If the formula of a compound is known (e.g. CytyCl), and you find the masses of the component atoms from the Periodic Table IC = 12 011g, H = 1 0079, & Cl = 35.455g). HOW would you find the molar mass of the compound?
    - b. Calculate the molar mass of one mole of CultuCl.
    - c. Calculate the mass of 5 miles of CJHJCI

#### E - Percent Composition of Compounds

- Provide the word equation describing how you solve the mass percent for a given element in a compound.
- 9) Penicillin is an important antibiotic that has the chemical formula C<sub>w</sub>H<sub>w</sub>N<sub>y</sub>SO<sub>w</sub>. Calculate the mass
- percent of EACH element in this compound.

#### F - Empirical Formulas

- 10) Consider the following:
- Taking in to account the formula for glucose, C<sub>a</sub>H<sub>ay</sub>O<sub>a</sub>, what information can be gethered about this compound?
  - b. What ratio of CH O is present when considering the formula above?
  - c. Now take in to account the following Formulas: CuHuOu, CuHuOu, and CuHuOu. What do these formulas have in common with glucose?
- 11) Using the word "ratio" in the descriptions, explain what the difference is between empirical and

# **Skill Practice 3answer Key Moles And Formulas**

**DJ Losen** 

#### **Skill Practice 3answer Key Moles And Formulas:**

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