

KEY

Molarity Practice

Name: _____

Date: _____

1. Calculate the molarity of the following solutions:

a) 45 g of Na_2SO_4 in 150 mL of solution.

$$(45\text{g}) \left(\frac{\text{mol}}{142.1\text{g}} \right) \left(\frac{1}{0.150\text{L}} \right) = \boxed{2.1\text{ M}}$$

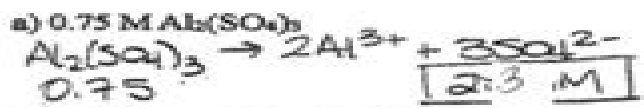
b) 24.6 g of $(\text{NH}_4)_2\text{CO}_3$ in 75 mL of solution.

$$(24.6\text{g}) \left(\frac{\text{mol}}{96.0\text{g}} \right) \left(\frac{1}{0.075\text{L}} \right) = \boxed{3.4\text{ M}}$$

c) 73.1 g of $\text{Ca}(\text{NO}_3)_2$ in 125 mL of solution.

$$(73.1\text{g}) \left(\frac{\text{mol}}{164.1\text{g}} \right) \left(\frac{1}{0.125\text{L}} \right) = \boxed{3.6\text{ M}}$$

2. What is the concentration of sulphate ions, SO_4^{2-} , in each of the following?



3. What is the molarity of chlorine ions in solution when 47 g of AlCl_3 is dissolved in a 210 mL of solution?

$$(47\text{g}) \left(\frac{\text{mol}}{133.5\text{g}} \right) \left(\frac{1}{0.210\text{L}} \right) = 1.68\text{ M}$$

$$[\text{Cl}^-] = 3(1.68\text{ M}) = \boxed{5.0\text{ M}}$$

4. Which of the following solutions has the highest concentration? Prove using calculations.

A) 12.5 g of CaCl_2 in 40 mL of solution

$$(12.5\text{g}) \left(\frac{\text{mol}}{111.1\text{g}} \right) \left(\frac{1}{0.040\text{L}} \right) = 2.82\text{ M}$$

B) 20.9 g of MgI_2 in 35 mL

$$(20.9\text{g}) \left(\frac{\text{mol}}{278.1\text{g}} \right) \left(\frac{1}{0.035\text{L}} \right) = 2.2\text{ M}$$

5. How many grams of salt (NaCl) need to be dissolved in 300 mL of solution to give you a solution that has a concentration of 1.2 M? (Hint: you need to work backwards on this one. You are given the molarity and the liters, so find the moles and convert to grams.)

$$\left(\frac{1.2\text{ mol}}{\text{L}} \right) (0.300\text{L}) = 0.36\text{ mol}$$

$$\text{m} = (0.36\text{ mol}) \left(\frac{58.5\text{g}}{\text{L}} \right) = \boxed{21.1\text{ g}}$$

Skill Practice 53 Answers Chemistry

Murray Shukyn, Dale E. Shuttleworth

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Ebbing Darrell D. Ebbing,1993 **Student Solutions Manual** Darrell Ebbing,2005-03-17 Provides solutions to odd numbered Practice Problems General Problems and Cumulative Skills Problems plus answers to Review Questions

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The academy, 1894 **An Introduction to Chemistry** Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

The Academy and Literature Charles Edward Cutts Birch Appleton, Charles Edward Doble, James Sutherland Cotton, Charles Lewis Hind, William Teignmouth Shore, Alfred Bruce Douglas, Ellis Ashmead-Bartlett, Thomas William Hodgson Crosland, 1894 [Chemistry Class XII For Madhya Pradesh Board by Dr. S C Rastogi, Er. Meera Goyal](#) Dr. S C Rastogi, Er. Meera Goyal, 2020-06-17 Syllabus Unit I Solid State Unit II Solutions Unit III Electrochemistry Unit IV Chemical Kinetics Unit V Surface Chemistry Unit VI General Principles and Processes of Isolation of Elements Unit VII p Block Elements Unit VIII d and f Block Elements Unit IX Coordination Compounds Unit X Haloalkanes and Haloarenes Unit XI Alcohols Phenols and Ethers Unit XII Aldehydes Ketones and Carboxylic Acids Unit XIII Organic Compounds Containing Nitrogen Unit XIV Biomolecules Unit XV Polymers Unit XV Polymers Unit XVI Chemistry in Everyday Life Content 1 Solid State 2 Solutions 3 Electro Chemistry 4 Chemical Kinetics 5 Surface Chemistry 6 General Principles And Processes Of Isolation Of Elements 7 P Block Elements 8 D And F Block Elements 9 Coordination Compounds And Organometallics 10 Haloalkanes And Haloarenes 11 Alcohols Phenols And Ethers 12 Aldehydes Ketones And Carboxylic Acids 13 Organic Compounds Containing Nitrogen 14 Biomolecules 15 Polymers 16 Chemistry In Everyday Life Appendix 1 Important Name Reactions And Process 2 Some Important Organic Conversions 3 Some Important Distinctions

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technology in the world a wondrous parade of machines and techniques such as PET scans MRI angioplasty endoscopy bypasses organ transplants and much more besides And yet writes Dr Eric Cassell what started out early in the century as the exciting conquest of disease has evolved into an overly expensive over technologized uncaring medicine poorly suited to the health care needs of a society marked by an aging population and a predominance of chronic diseases In *Doctoring The Nature of Primary Care Medicine* Dr Cassell shows convincingly how much better fitted advanced concepts of primary care medicine are to America s health care needs He offers valuable insights into how primary care physicians can be better trained to meet the needs of their patients both well and sick and to keep these patients as the focus of their practice Modern medical training arose at a time when medical science was in ascendancy Cassell notes Thus the ideals of science objectivity rationality became the ideals of medicine and disease the target of most medical research became the logical focus of medical practice When clinicians treat a patient with pneumonia they are apt to be thinking about pneumonia in general which is how they learn about the disease rather than this person s pneumonia This objective rational approach has its value but when it dominates a physician s approach to medicine it can create problems For instance treating chronic disease such as rheumatoid arthritis diabetes stroke emphysema and congestive heart failure is not simply a matter of medical knowledge for it demands a great deal of effort by the patients themselves they have to keep their doctor appointments take their medication do their exercises stop smoking The patient thus has a profound effect on the course of the disease and so for a physician to succeed he or she must also be familiar with the patient s motivations values concerns and relationship with the doctor Many doctors eventually figure out how to put the patient at the center of their practice but they should learn to do this at the training level not haphazardly over time To that end the training of primary care physicians must recognize a distinction between doctoring itself and the medical science on which it is based and should try to produce doctors who rely on both their scientific and subjective assessments of their patients overall needs There must be a return to careful observational and physical examination skills and finely tuned history taking and communication skills Cassell also advocates the need to teach the behavior of both sick and well persons evaluation of data from clinical epidemiology decision making skills and preventive medicine as well as actively teaching how to make technology the servant rather than the master and offers practical tips for instruction both in the classroom and in practice Most important *Doctoring* argues convincingly that primary care medicine should become a central focus of America s health care system not merely a cost saving measure as envisioned by managed care organizations Indeed Cassell shows that the primary care physician can fulfill a unique role in the medical community and a vital role in society in general He shows that primary care medicine is not a retreat from scientific medicine but the natural next step for medicine to take in the coming century

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