

(Actual Test total points = 72)

1.12a¹ One atom of an element would most likely weigh closest to 2 grams:

- a) 5×10^{23} b) 1×10^{23} c) 11×10^2 d) 0.1×10^{-23} e) 0.001×10^{-23}

b¹ The number of atoms in 1.0 moles of water is closest to:

- a) 3×10^{23} b) 1×10^{23} c) 11×10^2 d) 0.1×10^{-23} e) 0.1×10^{-26}

2.3¹ Why do chemists need to use the mole unit?2.4² Identify the reaction type for each equation below:4.1³ Be able to calculate molar mass. SHOW ALL WORK; LABEL WITH CORRECT UNITS!

a. Na

b. Fe

c. $(\text{CH}_3)_2\text{CHOH}$ d. F_2 e. P_4 f. $\text{Fe}(\text{OH})_3$ 4.4⁴ Balance the following equations:

4.6⁵ Consider the reaction: $\text{L} + \text{M} \rightarrow \text{X} + \text{Z}$. 57.0 g of L and 48.0 g of M are placed in a flask. After they are done reacting, L is all used up, 12.0 g of M remain, and 79.0 g of X were produced. What mass of Z must have been produced during the reaction? The data is summarized in the table below.

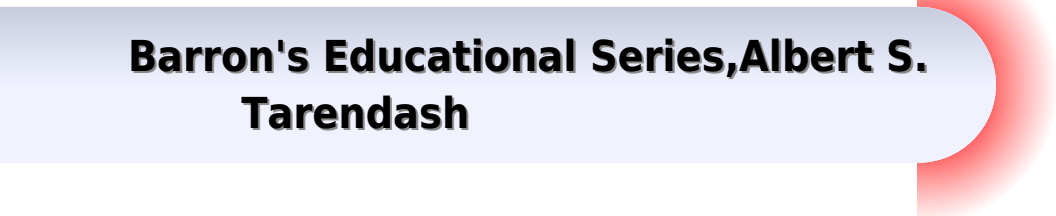
	L	M	X	Z
Mass before reaction	57.0 g	48.0 g	0 g	0 g
Mass after reaction	0 g	12.0 g	79.0 g	???

Answer: _____

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