

# STOICHIOMETRY NOTES WITH LIMITING REACTANT AND PERCENT YIELD

## STOICHIOMETRY

Stoichiometry is the relationship between the masses of the reactants and products of chemical reactions.

- Stoichiometry is possible because of the Law of Conservation of Mass
- The law of conservation of mass states that the amount of matter before a reaction occurs is equal to the amount of matter present after the reaction occurs.
- Thanks to this idea, we can build ratios between reactants and product that allow us to calculate amounts of each

## Using Ratios

Reactants and products are related through stoichiometry

Each coefficient in the balanced equation can be used to define a ratio between the individual reactants and products of the chemical reaction for the given balanced formula and division of divide

Consider the reaction:

$$2\text{NaCl} + \text{MgSO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{MgCl}_2$$

We can write down ratios:

|  |   |
|--|---|
| NaCl to NaCl                           | $\frac{2\text{NaCl}}{2\text{NaCl}}$ and $\frac{2\text{NaCl}}{2\text{NaCl}}$         |
| NaCl to NaCl                           | $\frac{2\text{NaCl}}{2\text{NaCl}}$ and $\frac{2\text{NaCl}}{2\text{NaCl}}$         |
| NaCl to NaCl                           | $\frac{2\text{NaCl}}{2\text{NaCl}}$ and $\frac{2\text{NaCl}}{2\text{NaCl}}$         |
| MgSO <sub>4</sub> to MgSO <sub>4</sub> | $\frac{1\text{MgSO}_4}{1\text{MgSO}_4}$ and $\frac{1\text{MgSO}_4}{1\text{MgSO}_4}$ |
| NaCl to NaCl                           | $\frac{2\text{NaCl}}{2\text{NaCl}}$ and $\frac{2\text{NaCl}}{2\text{NaCl}}$         |
| NaCl to NaCl                           | $\frac{2\text{NaCl}}{2\text{NaCl}}$ and $\frac{2\text{NaCl}}{2\text{NaCl}}$         |

## Example: Finding the Limiting Reactant in the reaction:

$$2\text{NaCl(aq)} + \text{MgSO}_4\text{(aq)} \rightarrow \text{Na}_2\text{SO}_4\text{(aq)} + \text{MgCl}_2\text{(aq)}$$

5.37 grams of NaCl react with 5.63 grams of Mg. What is the limiting reactant and what mass of  $\text{MgSO}_4$  can be formed?

How many moles of NaCl react?  $\frac{5.37 \text{ g NaCl}}{58.44 \text{ g/mol}} = 0.0919 \text{ mol NaCl}$

How many moles of Mg react?  $\frac{5.63 \text{ g Mg}}{24.31 \text{ g/mol}} = 0.232 \text{ mol Mg}$

How many moles of  $\text{MgSO}_4$  can be formed from the NaCl?  $0.0919 \text{ mol NaCl} \times \frac{1 \text{ mol MgSO}_4}{2 \text{ mol NaCl}} = 0.0459 \text{ mol MgSO}_4$

How many moles of  $\text{MgSO}_4$  can be formed from the Mg?  $0.232 \text{ mol Mg} \times \frac{1 \text{ mol MgSO}_4}{1 \text{ mol Mg}} = 0.232 \text{ mol MgSO}_4$

0.0459 moles of  $\text{MgSO}_4$  will form and NaCl is the limiting reactant

## Using Mole Ratios

From the last page, we looked at the reaction:

$$2\text{NaCl(aq)} + \text{MgSO}_4\text{(aq)} \rightarrow \text{Na}_2\text{SO}_4\text{(aq)} + \text{MgCl}_2\text{(aq)}$$

We also defined ratios of moles of each substance using the coefficients of the balanced equation.

How do we use the ratios we defined?

How many moles of Mg can be produced from 4 moles of NaCl in the reaction above?

We observe the ratio of moles of Mg to moles of NaCl, remembering to be the given number of moles of NaCl. By writing the ratio 4:1 in the denominator, we observe the ratio moles of NaCl and are left with moles of Mg.

$$4 \text{ moles NaCl} \times \frac{1 \text{ mol Mg}}{2 \text{ mol NaCl}} = 2 \text{ mol Mg}$$

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**Konstantinos Papadopoulos**



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