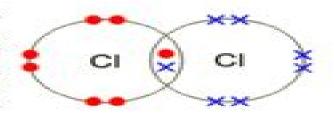
#### Covalent bonding

A non-metal metal and another non-metal can bond together by sharing electrons so that both electrons can achieve a full outer shell. This is called covalent bonding. To represent the atoms involved in the bonding we draw crosses for the electrons on one of the atoms and dots for the electrons on the other atoms as shown here.



#### Instructions

NB: It may be easiest to draw each atom out individually first to work out how many electrons will be shared

- 1. Draw the circles (outer shells) and overlap them (like a Venn diagram)
- 2. Write the symbols of the atoms involved in the middle of each circle
- Draw the electrons for one atom as dots (o) and electrons for the other atom as crosses (x).
- 4. Any shared electrons are drawn in the overlapping section.
- Count up the electrons for EACH atom and check that they each now have a full outer shell (2 electrons for the first shell and 8 for the next 2 shells)

TASK: Draw dot and cross diagrams for the following covalent compounds.

CH4	
CO2	.2
H <sub>2</sub> O	
NH₃	3
	HzO

Challenge! - Oz

# **The Nature Of Covalent Bonding 8 2 Worksheet**

Samuel N. Namowitz, Nancy E. Spaulding

The Nature Of Covalent Bonding 8 2 Worksheet:

Unveiling the Magic of Words: A Overview of "The Nature Of Covalent Bonding 8 2 Worksheet"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their power to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "**The Nature Of Covalent Bonding 8 2 Worksheet**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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