# Chemistry 105 Chapter 3 Review Problems

d) plant life will be exterminated

e) The CO2 will remove ozone from the stratosphere

 How many molecules are present in 180 grams of water? (The atomic weights of H and O are 1 g/mol and 16 g/mol, respectively.) a)  $6.0 \times 10^{24}$ b) 10: c)  $5.6 \times 10^{-23}$ d)  $1.8 \times 10^{2}$ e)  $180 \times 10^{23}$ 2. The atomic mass of sulfur is given as 32.07 g/mol. What is the mass of an average atom of sulfur? a) 5.33 x 10<sup>-25</sup> g b) 32.07 g c) 32.07 x 10<sup>-28</sup> kg d)  $193 \times 10^{-22}$  g e)  $6.02 \times 10^{23}$  g What is the mass of C present in 60 g of C<sub>2</sub>H<sub>6</sub>? a) 12 g b) 24 g c) 48 g d) 60 g e) 6 g 4. What is the molar mass of C<sub>2</sub>H<sub>6</sub>O? a) 24 g/molb) 32 g/mol c) 7 g/mol d) 35 g/mol e) 46 g/mol 5. What is the correct percentage by mass of nitrogen in HCN? a) 33 % b) 14% c) 27% d) 52% e) 98% What is the chief danger that is anticipated if CO<sub>2</sub> levels get too high in the atmosphere? a) people will develop skin cancer from exposure to UV light b) the atmosphere will develop an unpleasant odor c) the temperature of the planet may climb by several degrees

# **Standardized Test Practice Chapter 3 Chemistry**

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#### **Table of Contents Standardized Test Practice Chapter 3 Chemistry**

- 1. Understanding the eBook Standardized Test Practice Chapter 3 Chemistry
  - The Rise of Digital Reading Standardized Test Practice Chapter 3 Chemistry
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Standardized Test Practice Chapter 3 Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Standardized Test Practice Chapter 3 Chemistry
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Standardized Test Practice Chapter 3 Chemistry
  - Personalized Recommendations
  - Standardized Test Practice Chapter 3 Chemistry User Reviews and Ratings
  - Standardized Test Practice Chapter 3 Chemistry and Bestseller Lists
- 5. Accessing Standardized Test Practice Chapter 3 Chemistry Free and Paid eBooks
  - Standardized Test Practice Chapter 3 Chemistry Public Domain eBooks
  - Standardized Test Practice Chapter 3 Chemistry eBook Subscription Services
  - Standardized Test Practice Chapter 3 Chemistry Budget-Friendly Options

- 6. Navigating Standardized Test Practice Chapter 3 Chemistry eBook Formats
  - o ePub, PDF, MOBI, and More
  - Standardized Test Practice Chapter 3 Chemistry Compatibility with Devices
  - Standardized Test Practice Chapter 3 Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Standardized Test Practice Chapter 3 Chemistry
  - Highlighting and Note-Taking Standardized Test Practice Chapter 3 Chemistry
  - Interactive Elements Standardized Test Practice Chapter 3 Chemistry
- 8. Staying Engaged with Standardized Test Practice Chapter 3 Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Standardized Test Practice Chapter 3 Chemistry
- 9. Balancing eBooks and Physical Books Standardized Test Practice Chapter 3 Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Standardized Test Practice Chapter 3 Chemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Standardized Test Practice Chapter 3 Chemistry
  - Setting Reading Goals Standardized Test Practice Chapter 3 Chemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Standardized Test Practice Chapter 3 Chemistry
  - Fact-Checking eBook Content of Standardized Test Practice Chapter 3 Chemistry
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements

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