

## Laboratory experiment 2

### Standardization of Silver nitrate solution by Sodium chloride

**Objective:** to determine the normality (molar concentration of the equivalent) of silver nitrate solution.

Standardization of  $\text{AgNO}_3$  solution is based on the reaction:



**Procedure.** Completely fill the burette with the solution of silver nitrate  $\text{AgNO}_3$  and remove the air from the tip by running out some of the liquid into an empty beaker. Make sure that the lower part of the meniscus is at the zero mark.

Using a pipette, transfer a 20.00 mL aliquot of 0.1000 N solution of sodium chloride ( $\text{NaCl}$ ) to an Erlenmeyer flask and add 2-4 drops of potassium chromate  $\text{K}_2\text{CrO}_4$  solution ( $w=5\%$ ). Titrate this solution with silver nitrate solution: slowly add the  $\text{AgNO}_3$  solution to the flask with sodium chloride solution while gently stirring the contents of the flask. Continue titrating until the color of precipitate changes from white to reddish-brown.

Record the final burette reading. Repeat the titration two more times to a reproducible result. Calculate the normality of the  $\text{AgNO}_3$  and correction factor.

$$N(\text{AgNO}_3) = \frac{N(\text{NaCl}) \cdot V(\text{NaCl})}{V(\text{AgNO}_3)}$$

$$K = \frac{N_{\text{mean}}}{N_{\text{mean}}} =$$

Table

No	Volume of sodium chloride, mL	Volume of silver nitrate, mL	The mean of volume ( $\text{AgNO}_3$ ), mL	Normality of silver nitrate, mol/L
1	20.00	21.00		
2	20.00	20.90		
3	20.00	20.90		

**Conclusion:** \_\_\_\_\_

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# Standardization Of Agno3 Lab Report

**RJ Shavelson**



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