

# AP Chemistry: Chapter 14 Practice Test

10. Given the following acids and  $K_a$  values:



$$1 \times 10^7$$

$$1.76 \times 10^{-5}$$

$$4.93 \times 10^{-10}$$

$$3.53 \times 10^{-4}$$

which shows the conjugate bases listed by increasing strength?



15. Given that the  $K_a$  for  $\text{HOCl}$  is  $3.5 \times 10^{-8}$ , calculate the  $K$  value for the reaction of  $\text{HOCl}$  with  $\text{OH}^-$ .

[A]  $3.5 \times 10^{-22}$

[B]  $3.5 \times 10^6$

[C]  $2.9 \times 10^{-7}$

[D]  $3.5 \times 10^{-8}$

[E] none of these

32. Acetic acid ( $\text{HC}_2\text{H}_3\text{O}_2$ ) is a weak acid ( $K_a = 1.8 \times 10^{-5}$ ). Calculate the pH of a 17.6 M  $\text{HC}_2\text{H}_3\text{O}_2$  solution.

[A] 4.3

[B] 6.4

[C] 1.7

[D] 0.97

[E] 7.4

52. What concentration of acetic acid ( $K_a = 1.80 \times 10^{-5}$ ) has the same pH as that of  $5.00 \times 10^{-3}$  M  $\text{HCl}$ ?

[A]  $5.00 \times 10^{-3}$  M

[B]  $3.60 \times 10^{-3}$  M

[C] 1.39 M

[D] 5.00 M

[E] none of these

Consider the reaction  $\text{HOCl} + \text{F}^- \rightleftharpoons \text{HF} + \text{OCl}^-$

# Standardized Test Practice Chapter 14 Chemistry Answers

**Clifford Lane**



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