

NAME: \_\_\_\_\_

CLASS: \_\_\_\_\_

DATE: \_\_\_\_\_

1. What is the percent by mass of oxygen in MgO?

- |                         |     |                         |     |
|-------------------------|-----|-------------------------|-----|
| <input type="radio"/> A | 50% | <input type="radio"/> B | 60% |
| <input type="radio"/> C | 20% | <input type="radio"/> D | 40% |

2. How many molecules are in 2.5 mol of NaCl?

- |                         |                |                         |                                 |
|-------------------------|----------------|-------------------------|---------------------------------|
| <input type="radio"/> A | 4.15 molecules | <input type="radio"/> B | $1.51 \times 10^{24}$ molecules |
| <input type="radio"/> C | 1.50 molecules | <input type="radio"/> D | $1.50 \times 10^{24}$ molecules |

3. A student calculated the number of moles of methane ( $\text{CH}_4$ ) present in a 100g sample of a gaseous mixture. If the student determined that there's 0.6 moles of methane present in the sample, what is the percent composition of methane in the sample?

- |                         |      |                         |       |
|-------------------------|------|-------------------------|-------|
| <input type="radio"/> A | 9.6% | <input type="radio"/> B | .096% |
| <input type="radio"/> C | .9%  | <input type="radio"/> D | 90%   |

4. A 15.67 g sample of a hydrate of magnesium carbonate was heated to drive off the water. The mass was reduced to 7.58 g. What is the formula of the hydrate?

- |                         |   |                         |   |
|-------------------------|---|-------------------------|---|
| <input type="radio"/> A | $\text{MgCO}_3 \cdot 5\text{H}_2\text{O}$ | <input type="radio"/> B | $\text{MgCO}_3 \cdot 4\text{H}_2\text{O}$ |
| <input type="radio"/> C | $\text{MgCO}_3 \cdot 6\text{H}_2\text{O}$ | <input type="radio"/> D | $\text{MgCO}_3 \cdot 1\text{H}_2\text{O}$ |

5. What is the empirical formula of a compound that is 81.82% carbon and 18.18% hydrogen?

- |                         |                           |                         |                        |
|-------------------------|---------------------------|-------------------------|------------------------|
| <input type="radio"/> A | $\text{C}_4\text{H}_{10}$ | <input type="radio"/> B | $\text{C}_2\text{H}_2$ |
| <input type="radio"/> C | $\text{CH}_4$             | <input type="radio"/> D | $\text{C}_3\text{H}_8$ |

6.  $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$

How many moles of oxygen are produced when 6.7 moles of  $\text{KClO}_3$  decompose completely?

# Star Test 10th Grade Chemistry

**SJ Ball**



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