

## Experiment 9

### Acid-Base Titration

#### OBJECTIVES:

- To learn basic chemical characteristics of acids and bases.
- To learn how to perform titrations with precision.
- To understand how titrations can be used in chemical analysis.
- To prepare and standardize an NaOH solution.
- To use your standardized NaOH to determine the molarity of acetic acid in vinegar.

#### INTRODUCTION:

All Freshman Chemistry labs include at least one **titration** experiment. Titration is an old and time-honored method of **quantitative analysis**.<sup>1</sup> In every titration, we have some reaction occurring between some pair of reactants, A & B.

##### *General*



##### *Reaction*

A titration reaction usually falls into one of the three following categories:

(1) A Precipitation Reaction.



(2) An Acid-Base Reaction.

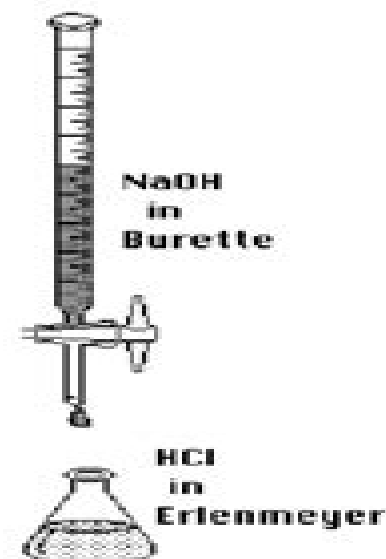


(3) An Oxidation-Reduction Reaction.



In today's experiment the reaction will be of type (2)—an acid-base neutralization reaction. As always in a titration, we place one reactant (a solution of NaOH) in a Burette and the other

<sup>1</sup> Quantitative Analysis techniques are designed to answer quantitative questions concerning how much of something is present in a sample. Quantitative Analysis techniques are to be distinguished from **Qualitative Analysis** techniques that are designed to answer qualitative questions concerning what substances are present in a sample.



# Virtual Acid Base Ph Titration Packet Answers

**Gregory K. McMillan, Robert A.  
Cameron**



## Virtual Acid Base Ph Titration Packet Answers:

*Instructor's Manual* Brian F. Woodfield, Matthew C. Asplunc, 2006 NEW Click here to visit the Virtual ChemLab Frequently Asked Questions FAQ document This Instructor's Lab Manual Workbook is similar to the Student Lab Manual Workbook and additionally contains an overview of the full capabilities of the Site License version of Virtual ChemLab installation instructions and the answers for the laboratory assignments provided in the student laboratory workbook This product is available within Virtual ChemLab General Chemistry Instructor Lab Manual Workbook and Student CD Combo Package v2 5 0 13 228010 8 Valuepack and or should be ordered in conjunction with Virtual ChemLab General Chemistry Instructor Site License CD v2 5 0 13 185749 5     Ph Scale And Acid-base Titrations/standardization Postma Roberts Hollenberg, James M. Postma, Julian Roberts, J. Leland Hollenberg, 2000-01-15     PH, Buffers and Acid-base Titrations LaMotte Chemical Products Co. Educational Products Division, 197?     Aqueous Acid-base Equilibria and Titrations Robert De Levie, 1999 This book will give students a thorough grounding in pH and associated equilibria material absolutely fundamental to the understanding of many aspects of chemistry It is in addition a fresh and modern approach to a topic all too often taught in an outmoded way This book uses new theoretical developments which have led to more generalized approaches to equilibrium problems these approaches are often simpler than the approximations which they replace Acid base problems are readily addressed in terms of the proton condition a convenient amalgam of the mass and charge constraints of the chemical system considered The graphical approach of Bjerrum Hagg and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species involved in chemical equilibria Based on these concentrations the proton condition can usually be simplified often leading directly to the value of the pH In the description of acid base titrations a general master equation is developed It provides a continuous and complete description of the entire titration curve which can then be used for computer based comparison with experimental data Graphical estimates of the steepness of titration curves are also developed from which the practicality of a given titration can be anticipated Activity effects are described in detail including their effect on titration curves The discussion emphasizes the distinction between equilibrium constants and electrometric pH measurements which are subject to activity corrections and balance equations and spectroscopic pH measurements which are not Finally an entire chapter is devoted to what the pH meter measures and to the experimental and theoretical uncertainties involved     *PH Measurement and Titration* G. Mattock, 1961     *Advanced Ph Measurement and Control* Gregory K. McMillan, Robert A. Cameron, 2004-10-30     An On-line Acid-base Titration Applet in the Generic Tutorial System for the Sciences Project Thomas Lee Gummo, 2002 The purpose of this Master's Project was to develop an Acid Base Titration Simulator It was also to be a part of the California State University San Bernardino's GTSS Generic Tutorial System for the Sciences project The main benefit is that students will be able to conduct titration experiments over the Internet without being in the laboratory and without costly equipment or dangerous chemicals

Instructors at the high school and college level can demonstrate the key principles of titration ACID-BASE EQUILIBRIA

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