

Atomic Emission Spectra Lab with Key

Teacher Notes

This lab is designed to be a fun, hands-on activity for students. It is a great way to introduce the concept of atomic emission spectra and the Bohr model of the atom. The lab is designed to be a fun, hands-on activity for students. It is a great way to introduce the concept of atomic emission spectra and the Bohr model of the atom.

Use this lab as a demonstration or as a student activity. The lab is designed to be a fun, hands-on activity for students. It is a great way to introduce the concept of atomic emission spectra and the Bohr model of the atom.

Be careful handling the gas tubes. Only touch them off the burner and not the burner itself. They can get very hot. Do not touch them if they are glowing. Do not touch them if they are glowing.

You can also use this lab as a student activity. It is a great way to introduce the concept of atomic emission spectra and the Bohr model of the atom.

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1. The atomic emission spectra of hydrogen, helium, and neon are shown below. The emission spectra of these elements are shown below. The emission spectra of these elements are shown below.

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Gas Sample	Observed Emission Lines	Observed Wavelengths (nm)
Hydrogen (H ₂)	4 visible lines (violet, blue-violet, yellow-green, red)	410.1, 434.0, 486.1, 656.3
Helium (He)	5 visible lines (violet, blue-violet, yellow-green, red)	402.6, 437.9, 504.8, 587.6, 667.8
Neon (Ne)	5 visible lines (violet, blue-violet, yellow-green, red)	444.1, 505.0, 588.2, 640.0, 780.0

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Virtual Chem Lab 5 Atomic Emission Spectra

Steve Suib

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name 'Steve Suib'.

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