

# Stoichiometry

## Section 11.1 Defining Stoichiometry

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### Practice Problems

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1. Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is observed.



1 molecule  $\text{N}_2$  + 3 molecules  $\text{H}_2$   
 $\rightarrow$  2 molecules  $\text{NH}_3$

1 mole  $\text{N}_2$  + 3 moles  $\text{H}_2 \rightarrow$  2 moles  $\text{NH}_3$

Mass:

$\text{N}_2$ : (2 mol)(14.007 g/mol) = 28.014 g

$3\text{H}_2$ : (6 mol)(1.008 g/mol) = 6.048 g

$2\text{NH}_3$ : (2 mol)(14.007 g/mol) +  
 (6 mol)(1.008 g/mol) = 34.062 g

28.014 g  $\text{N}_2$  + 6.048 g  $\text{H}_2 \rightarrow$  34.062 g  $\text{NH}_3$

34.062 g reactants = 34.062 g products



1 molecule  $\text{HCl}$  + 1 formula unit  $\text{KOH} \rightarrow$   
 1 formula unit  $\text{KCl}$  + 1 molecule  $\text{H}_2\text{O}$

1 mole  $\text{HCl}$  + 1 mole  $\text{KOH} \rightarrow$   
 1 mole  $\text{KCl}$  + 1 mole  $\text{H}_2\text{O}$

Mass:

$\text{HCl}$ : (1 mol)(1.008 g/mol) +  
 (1 mol)(35.453 g/mol) = 36.461 g

$\text{KOH}$ : (1 mol)(39.098 g/mol) +  
 (1 mol)(15.999 g/mol) + (1 mol)(1.008 g/mol)  
 = 56.105 g

$\text{KCl}$ : (1 mol)(39.098 g/mol) +  
 (1 mol)(35.453 g/mol) = 74.551 g

$\text{H}_2\text{O}$ : (2 mol)(1.008 g/mol) +  
 (1 mol)(15.999 g/mol) = 18.015 g

36.461 g  $\text{HCl}$  + 56.105 g  $\text{KOH} \rightarrow$   
 74.551 g  $\text{KCl}$  + 18.015 g  $\text{H}_2\text{O}$

92.566 g reactants = 92.566 g products



2 atoms  $\text{Mg}$  + 1 molecule  $\text{O}_2$   
 $\rightarrow$  2 formula units  $\text{MgO}$

2 moles  $\text{Mg}$  + 1 mole  $\text{O}_2 \rightarrow$  2 moles  $\text{MgO}$

Mass:

$2\text{Mg}$ : (2 mol)(24.305 g/mol) = 48.610 g

$\text{O}_2$ : (2 mol)(15.999 g/mol) = 31.998 g

$2\text{MgO}$ : (2 mol)(24.305 g/mol) +  
 (2 mol)(15.999 g/mol) = 80.608 g

48.610 g  $\text{Mg}$  + 31.998 g  $\text{O}_2 \rightarrow$  80.608 g  $\text{MgO}$

80.608 g reactants = 80.608 g products

2. **Challenge** For each of the following, balance the chemical equation; interpret the equation in terms of particles, moles, and mass; and show that the law of conservation of mass is observed.



$2\text{Na}(\text{s}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow 2\text{NaOH}(\text{aq}) + \text{H}_2(\text{g})$

2 atoms  $\text{Na}$  + 2 molecules  $\text{H}_2\text{O} \rightarrow$   
 2 formula units  $\text{NaOH}$  + 1 molecule  $\text{H}_2$

2 mol  $\text{Na}$  + 2 mol  $\text{H}_2\text{O} \rightarrow$   
 2 mol  $\text{NaOH}$  + 1 mol  $\text{H}_2$

Mass:

$2\text{Na}$ : (2 mol)(22.990 g/mol) = 45.980 g

$2\text{H}_2\text{O}$ : (4 mol)(1.008 g/mol) +  
 (2 mol)(15.999 g/mol) = 36.030 g

$2\text{NaOH}$ : (2 mol)(22.990 g/mol) +  
 (2 mol)(15.999 g/mol) + (2 mol)(1.008 g/mol)  
 = 79.994 g

$\text{H}_2$ : (2 mol)(1.008 g/mol) = 2.016 g

45.980 g  $\text{Na}$  + 36.030 g  $\text{H}_2\text{O} \rightarrow$   
 79.994 g  $\text{NaOH}$  + 2.016 g  $\text{H}_2$

82.01 g reactants = 82.01 g products

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