

# CHEMICAL EQUILIBRIUM

➤  $K_c = \frac{[B]}{[A]} = \text{mol L}^{-1}$

➤  $K_p = \frac{P_B}{P_A} = \text{Partial Pressure}$

➤  $K_x = \frac{X_B}{X_A} = \text{mole fraction}$

➤  $K_p = K_c(RT)^{\Delta n_g}$  where  $\Delta n_g = n_p - n_r$

➤ Predicting the extent of reaction :

- $K_c > 10^3$  [Forward reaction is favoured.]
- $K_c < 10^{-3}$  [Reverse reaction is favoured.]
- $10^{-3} < K_c < 10^3$  [Both reactants and products are present in equilibrium]

➤ Free Energy Change ( $\Delta G$ )

- If  $\Delta G = 0$  then reversible reaction would be in equilibrium,  $K_c = 1$
- If  $\Delta G = (+)$  ve then equilibrium will be displaced in backward direction;  $K_c < 1$
- If  $\Delta G = (-)$  ve then equilibrium will shift in forward direction;  $K_c > 1$

➤ (a)  $K_c$  unit  $\rightarrow (\text{moles/lit})^{\Delta n_g}$

➤  $\Delta G$  or a reaction under any set of conditions is related to its value under standard conditions, i.e.  $\Delta G^\circ$  by the equation

$$\Delta G = \Delta G^\circ + 2.303 RT \log Q$$

➤ Under equilibrium conditions, for same number of moles of reactants and products

$$Q = K_p = K_c = K \text{ and } \Delta G = 0$$

$\therefore$

$$\Delta G^\circ = -2.303 RT \log K$$

➤ We have replaced  $K_p$  by  $K$  called thermodynamic equilibrium constant

## Significance of $\Delta G^\circ$

Significance of  $\Delta G^\circ$  can be explained from the following points

(i) If  $\Delta G^\circ < 0$ ,  $\log K > 0$ ,  $\Rightarrow K > 1$

Hence, reaction is spontaneous in forward direction.

(ii) Hence  $\Delta G^\circ < 0$ ,  $\log K > 0$ ,  $\Rightarrow K < 1$

Hence, reaction is non-spontaneous or a reaction proceeds in the forward direction to such a small extent that a very small amount of the product is formed.



# Supplemental Problems Answer Sheet Chemical Equilibrium

**Christian G. Meyer**



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