

## ELECTROCHEMISTRY

### A. Matching

Match each term in Column B to the correct description in Column A.

| Column A                                                                                                        | Column B       |
|-----------------------------------------------------------------------------------------------------------------|----------------|
| 1. a conductor in a circuit that carries electrons to or from a substance other than a metal                    | ✓ anode        |
| 2. one part of a voltaic cell, in which either oxidation or reduction occurs                                    | ✓ electrode    |
| 3. a voltaic cell in which the electrolyte is a paste                                                           | ✓ cathode      |
| 4. the electrode at which reduction occurs                                                                      | ✓ dry cell     |
| 5. a voltaic cell in which a fuel undergoes oxidation and from which electrical energy is obtained continuously | ✓ battery      |
| 6. an electrochemical cell that is used to convert chemical energy into electrical energy                       | ✓ electrolysis |
| 7. a tube containing a conducting solution                                                                      | ✓ voltaic cell |
| 8. a group of cells that are connected together                                                                 | ✓ half-cell    |
| 9. the process in which electrical energy is used to bring about a chemical change                              | ✓ salt bridge  |
| 10. the electrode at which oxidation occurs                                                                     | ✓ fuel cell    |

### B. Multiple Choice

Choose the best answer and write its letter on the line.

11. In the electrolysis of brine, the substance produced at the cathode is  
a. chlorine. c. sodium.  
b. oxygen. d. hydrogen.
12. Which of the following describes a dry cell?  
a. It can be recharged many times.  
b. The graphite rod does not undergo reduction, even though it is the cathode.  
c. It contains concentrated sulfuric acid.  
d. all of the above.

# Study Guide For Content Mastery Electrochemistry

**Kaplan Test Prep**



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to *Journals in Education*, 1990      **Forthcoming Books** Rose Arny, 2002      *Compendium of Science Demonstration-related Research from 1918 to 2008* David Michael Majerich, Joseph S. Schmuckler, Kathleen Fadigan, 2008 This compendium highlights a ninety year history of science education focused on how demonstrations have been used for the express teaching and learning of science at the high school college and university levels It is presented in three chapters a Experimental Studies Comparing the Lecture Demonstration Method and the Individual Laboratory Method of Teaching Science 1918 1989 b Further Types of Demonstration related Comparison Studies 1958 2008 and c More Recent Demonstration related Non Comparative Studies 1980 2008 Organization of the research in this manner revealed that an interest in investigating this type of research by science educators had and continues to evolve This work was constructed from an intense and nearly exhaustive review of the research While the authors believe that it can be misconstrued that the scholarly literature and empirical research are one in the same we can show that it is necessary to make a distinction between the two sources of information Upon review of the scholarly literature and empirical research on the demonstration topics a salient outcome is that these two bodies of information are only loosely corroborated From this point on we use the terms literature and research when referring to each of these bodies of information To begin we describe studies comparing the effectiveness of the methods of lecture demonstration and the individual laboratory The very early attempts 1918 1964 to determine the superiority of the lecture demonstration method over the individual laboratory method of teaching science and vice versa remained unresolved through the mid 1960s Cunningham 1946 suggested several reasons for the inconclusive nature of the results reported by researchers of lecture demonstration and individual laboratory comparison studies After carefully reviewing the studies up to 1946 he noted that there were numerous variables that should have remained fixed for the duration of the experiments Some of the variables that confounded the results of those investigations included the uncontrolled variables related to a the teacher b the complexity of experiments and apparatus c the time spent on each method d the amount of science studied by students and e the performer of the demonstrations Even after exacting a procedure to control for the teacher variable reported by Cunningham 1946 Yager et al 1969 were able to show that students who participated in a discussion demonstration group or discussion laboratory group developed more skills than those students who received science instruction via a discussion only method The results of Yager and his predecessors suggested that neither the demonstration method nor the laboratory method for the teaching and learning of science was superior From 1958 to 2008 investigations comprised of multiple demonstration related comparison studies were also unable to confirm the superiority of the demonstration method for the teaching and learning of science as compared to other methods in each study For instance Oliver 1975 compared three methods of teaching high school biology lecture discussion a combination of lecture discussion and demonstration and demonstrations Initially although he determined that the lecture discussion method of teaching biology was noticeably superior in terms of biology content acquisition during the onset of the semester this result

was ephemeral and dissipated at the termination of the following semester Comparing a self paced instructional method and a teacher demonstration method of teaching college chemistry Eniaijeju 1983 concluded that students achievement scores were higher when they participated in the self paced program in

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