

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for tertiary hydroxyl groups



A Summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: ionic bond and covalent bond

- Ionic bonds are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron affinity).



- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Bonding in a Molecule



Bonding in a Molecule



There are two types of covalent bond

Definition of ionic bond



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds

- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- A single bond is a sigma bond.
- A double bond consists of one sigma and one pi bond.



pi bonds are stronger than sigma bonds

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds are divided into polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

A difference of 1.7 or higher indicates that the electrons are not shared (ionic bond) and a lower value indicates a covalent bond.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is defined as a vector pointing from the less electronegative to the more electronegative atom.
- The vector dipole moment is a vector pointing from the less electronegative to the more electronegative atom.
- $\mu = \delta \times r$ (where δ is the charge and r is the distance between the charges).
- μ (dipole moment) indicates the direction of the bond.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is calculated.

The polarity of the molecule is determined by the vector sum of the dipole moments. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipole moment is zero.

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.

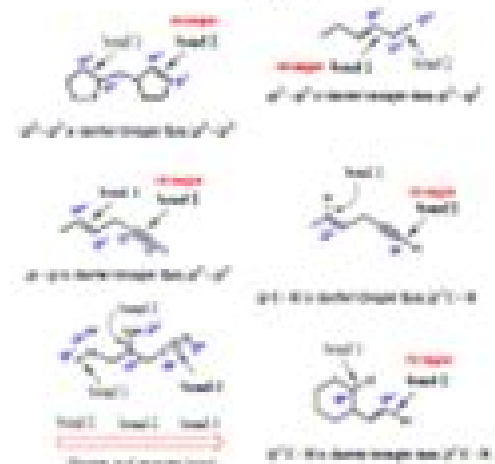


The bond length is the distance between the nuclei of the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.09 Å	1.08 Å	1.06 Å
Strength	98 kcal/mol	110 kcal/mol	120 kcal/mol

Higher s character increases bond strength

Example



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Table of Contents Study Guide For Organic Chemistry

1. Understanding the eBook Study Guide For Organic Chemistry
 - The Rise of Digital Reading Study Guide For Organic Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Organic Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
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 - Popular eBook Platforms
 - Features to Look for in an Study Guide For Organic Chemistry
 - User-Friendly Interface
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 - Creating a Diverse Reading Collection Study Guide For Organic Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide For Organic Chemistry
 - Setting Reading Goals Study Guide For Organic Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide For Organic Chemistry
 - Fact-Checking eBook Content of Study Guide For Organic Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

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- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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