

116. The rate of effusion of an unknown gas was measured and found to be 11.9 mL/min. Under identical conditions, the rate of effusion of pure oxygen (O_2) gas is 14.0 mL/min. Based on this information, the identity of the unknown gas could be:

A) F_2
B) NO
C) CO_2
D) C_2H_2
E) none of these

ANS: C PTS: 1 DIF: Moderate REF: 5.7
KEY: Chemistry / general chemistry / phases / gas / molecular speed / effusion
MSC: Quantitative

117. Calculate the ratio of the effusion rates of N_2 and N_2O .

A) 0.637
B) 1.57
C) 1.25
D) 0.788
E) 1.62

ANS: C PTS: 1 DIF: Moderate REF: 5.7
KEY: Chemistry / general chemistry / phases / gas / molecular speed / effusion
MSC: Conceptual

118. Hydrogen and chlorine gases react to form HCl . You and a friend are on opposite sides of a long hallway, you with H_2 and your friend with Cl_2 . You both want to form HCl in the middle of the room. Which of the following is true?

A) You should release the H_2 first.
B) Your friend should release the Cl_2 first.
C) You both should release the gases at the same time.
D) You need to know the length of the room to answer this question.
E) You need to know the temperature to answer this question.

ANS: B PTS: 1 DIF: Easy REF: 5.7
KEY: Chemistry / general chemistry / phases / gas / molecular speed / diffusion
MSC: Conceptual

119. Order the following in increasing rate of effusion:

F_2 , Cl_2 , NO , NO_2 , CH_4

A) $Cl_2 < NO_2 < F_2 < NO < CH_4$
B) $Cl_2 < F_2 < NO_2 < CH_4 < NO$
C) $CH_4 < NO_2 < NO < F_2 < Cl_2$
D) $CH_4 < NO < F_2 < NO_2 < Cl_2$
E) $F_2 < NO < Cl_2 < NO_2 < CH_4$

ANS: A PTS: 1 DIF: Easy REF: 5.7
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