

Modeled after AP 2013 #21:

- $$2 \text{C}_2\text{H}_6\text{g} + 7 \text{O}_2\text{g} \rightarrow 4 \text{CO}_2\text{g} + 6 \text{H}_2\text{Og}$$

In the reaction represented above, what is the total number of moles of reactants consumed when 2.00 mole of  $\text{CO}_2\text{g}$  is produced?

  - 0.50 mol
  - 1.50 mol
  - 4.50 mol**
  - 5.00 mol
- $$\text{C}_2\text{H}_6\text{g} + 5 \text{O}_2\text{g} \rightarrow 4 \text{CO}_2\text{g} + 6 \text{H}_2\text{Og}$$

In the reaction represented above, what is the total number of moles of reactants consumed when 2.00 mole of  $\text{CO}_2\text{g}$  is produced?

  - 0.50 mol
  - 1.00 mol
  - 3.50 mol**
  - 7.00 mol
- $$2 \text{C}_2\text{H}_6\text{g} + 7 \text{O}_2\text{g} \rightarrow 4 \text{CO}_2\text{g} + 6 \text{H}_2\text{Og}$$

In the reaction represented above, what is the total number of moles of products made when 3.50 mole of  $\text{O}_2\text{g}$  is consumed?

  - 2.00 mol
  - 3.00 mol
  - 5.00 mol**
  - 10.0 mol
- $$\text{C}_2\text{H}_6\text{g} + 5 \text{O}_2\text{g} \rightarrow 4 \text{CO}_2\text{g} + 6 \text{H}_2\text{Og}$$

In the reaction represented above, what is the total number of moles of products made when 3.50 mole of  $\text{O}_2\text{g}$  is consumed?

  - 1.00 mol
  - 2.00 mol**
  - 4.00 mol
  - 8.00 mol

Modeled after AP 2014g #48:

- $$4 \text{NH}_3\text{g} + 5 \text{O}_2\text{g} \rightarrow 4 \text{NOg} + 6 \text{H}_2\text{Og}$$

A 7.0 mol sample of  $\text{NH}_3\text{g}$  and a 10. mol sample of  $\text{O}_2\text{g}$  are placed in a previously evacuated container and react according to the equation above. How many moles of  $\text{NOg}$  are produced when one of the reactants has been totally consumed?

  - 2.5 mol
  - 7.0 mol**
  - 8.0 mol
  - 28 mol
- $$\text{C}_2\text{H}_6\text{g} + 3 \text{H}_2\text{g} \rightarrow 2 \text{CH}_4\text{g}$$

Samples containing 4.0 mol  $\text{C}_2\text{H}_6\text{g}$  and 5.0 mol  $\text{H}_2\text{g}$  are placed in a previously evacuated container and allowed to react according to the equation above. When one of the reactants is completely consumed, how much product is formed?

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