NUCLEAR CHEMISTRY STUDY GUIDE—CH. 21

- Write the symbol for uranium—238 in isotope format (with top and bottom #s).
- What does the top # represent?
- 3. What does the bottom # represent?
- 4. What particles are found in the nucleus?
- 5. Why does the strong nuclear force exist?
- 6. What is the electrostatic force?
- 7. What particles does the electrostatic force deal with?
- 8. Define nuclear reactions.
- Define nuclear fission and fusion.
- 10. Nuclear reactions usually produce ______amounts of energy, and isotopes become _____stable. The _____number and the _____number may also change.
- 11. What is transmutation?
- 12. Write the beta decay of radium-226.
- 13. Write the alpha decay of radium-226.
- 14. Why would the isotopes and the end of problems #13 and #14 still be unstable?
- 15. Define nuclear radiation.
- 16. List the 3 radioactive particles in order of INCREASING penetrating power.
- 17. What is radioactive decay?
- Alpha particles are basically _____nuclei, and beta particles are basically high energy _____.
- 19. What is half life?
- 20. What is radioactive dating?
- 21. What is cobalt-60 used for?
- 22. Does fission or fusion produce nuclei with heavier mass than the original mass?
- 23. Name 3 differences between chemical and nuclear reactions.
- 24. What is carbon-13 radioactive?
- Elements past atomic # ______are radioactive.
- 26. If 1/16 of a sample of a radioactive isotope Is left after nuclear decay, how long did this take if the ½ life is 4 days?
- 27. If the half life of a radioactive isotope is 4 days, how much of a 50 gram sampel is left after 12 days?

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Rachel S Tattersall

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