

Name _____ Date _____ Class _____

CHAPTER 15 **STUDY GUIDE FOR CONTENT MASTERY**

Solutions

Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

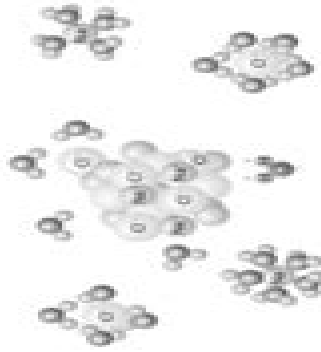
Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solvent	solvent

Air is air! (1) solution of oxygen gas dissolved in nitrogen gas. The oxygen in air is the (2) solute and nitrogen is the (3) solvent. Because oxygen gas dissolves in a solvent, oxygen gas is air's (4) soluble substance. A substance that does not dissolve is (5) insoluble. (6) Liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are (7) miscible. However, if one liquid is insoluble in another, the liquids are (8) immiscible.

Read about solvation in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



8. Hydration is solvation in which the solvent is water. What is solvation?

Solvation is the process of surrounding solute particles with solvent particles to form a solution.

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CHAPTER 15 **STUDY GUIDE FOR CONTENT MASTERY**

Section 15.1 continued

10. As sodium chloride dissolves in water, what happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

11. Explain the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

12. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.

13. List three ways that the rate of solvation may be increased.

stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write true or false.

true 14. The overall energy change that occurs when a solution forms is called the heat of solution.

false 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

false 16. Solvation continues as long as the solvation rate is less than the crystallization rate.

true 17. In a saturated solution, solvation and crystallization are in equilibrium.

true 18. Additional solute can be dissolved in an unsaturated solution.

true 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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Study Guide For Content Master Chemistry Key

Jenny Frost, Tony Turner



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Life William K. Purves, 2004 New edition of a text presenting underlying concepts and showing their relevance to medical, agricultural and environmental issues. Seven chapters discuss the cell, information and heredity, evolutionary process, the evolution of diversity, the biology of flowering plants and of animals and ecology and biogeography. Topics are linked by themes such as evolution, the experimental foundations of knowledge, the flow of energy in the living world, the application and influence of molecular techniques and human health considerations. Includes a CD-ROM which covers some of the subject matter and introduces and illustrates 1 700 plus key terms and concepts. Annotation copyrighted by Book News Inc. Portland, OR

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Midwest *CBSE Class 12 Mastering Organic Reactions MCQ With Answers* Priti Singhal, 2024-11-11 The primary objective of this book is to serve as a comprehensive guide for students, educators and researchers by focusing on reaction mechanisms, practical applications and problem solving techniques. Organic chemistry is not just about memorizing equations and formulas; it is about understanding how molecules interact, change and influence each other under different conditions. With that in mind, this book emphasizes the logic and patterns behind organic reactions, making it easier for readers to apply concepts across a variety of scenarios. Each chapter of this book builds upon foundational knowledge, ensuring a progressive learning experience. From nucleophilic substitutions to pericyclic reactions and from oxidation-reduction mechanisms to named reactions, we cover both fundamental and advanced topics to cater to students at all levels. Real world examples have been integrated throughout the chapters to show how organic reactions play essential roles in pharmaceuticals, biochemistry, agriculture and environmental science. This approach bridges the gap between theory and practical applications, helping readers appreciate the relevance of organic chemistry in daily life.

Standards Matter Katherine Scheidler, 2017-05-01 While critics rail against Common Core State Standards for national school learning guides, few know exactly what these Standards are and fewer can assess if these Standards are a positive step for education. Standards are simply the high level

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