

## AP Chemistry Chapter 4 Answers – Zumdahl

4.36

The solubility rules referenced in the following answers are from Table 4.1 of the text. The phrase “slightly soluble” is interpreted to mean insoluble and the phrase “marginally soluble” is interpreted to mean soluble.

- a. Soluble (Rule 3)
- b. Soluble (Rule 1)
- c. Insoluble (Rule 4)
- d. Soluble (Rules 2 and 3)
- e. Insoluble (Rule 6)
- f. Insoluble (Rule 5)
- g. Insoluble (Rule 6)
- h. Soluble (Rule 2)

4.38

Use Table 4.1 to predict the solubility of the possible products.

- a. Possible products =  $\text{Hg}_2\text{SO}_4$  and  $\text{Cu}(\text{NO}_3)_2$ ; precipitate =  $\text{Hg}_2\text{SO}_4$
- b. Possible products =  $\text{NiCl}_2$  and  $\text{Ca}(\text{NO}_3)_2$ ; Both salts are soluble so no precipitate forms.
- c. Possible products =  $\text{KI}$  and  $\text{MgCO}_3$ ; precipitate =  $\text{MgCO}_3$
- d. Possible products =  $\text{NaBr}$  and  $\text{Al}_2(\text{CrO}_4)_3$ ; precipitate =  $\text{Al}_2(\text{CrO}_4)_3$

4.44

- a.  $\text{CrCl}_3(\text{aq}) + 3\text{NaOH}(\text{aq}) \rightarrow \text{Cr}(\text{OH})_3(\text{s}) + 3\text{NaCl}(\text{aq})$   
 $\text{Cr}^{3+}(\text{aq}) + 3\text{OH}^-(\text{aq}) \rightarrow \text{Cr}(\text{OH})_3(\text{s})$
- b.  $2\text{AgNO}_3(\text{aq}) + (\text{NH}_4)_2\text{CO}_3(\text{aq}) \rightarrow \text{Ag}_2\text{CO}_3(\text{s}) + 2\text{NH}_4\text{NO}_3(\text{aq})$   
 $2\text{Ag}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{Ag}_2\text{CO}_3(\text{s})$
- c.  $\text{CuSO}_4(\text{aq}) + \text{Hg}_2(\text{NO}_3)_2(\text{aq}) \rightarrow \text{Cu}(\text{NO}_3)_2(\text{aq}) + \text{Hg}_2\text{SO}_4(\text{s})$   
 $\text{Hg}_2^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) \rightarrow \text{Hg}_2\text{SO}_4(\text{s})$
- d. No reaction occurs because all possible products ( $\text{SrI}_2$  and  $\text{KNO}_3$ ) are soluble.

4.46

Because no precipitates formed upon addition of  $\text{NaCl}$  or  $\text{Na}_2\text{SO}_4$ , we can conclude that  $\text{Hg}_2^{2+}$  and  $\text{Ba}^{2+}$  are not present in the sample since  $\text{Hg}_2\text{Cl}_2$  and  $\text{BaSO}_4$  are insoluble salts. However,  $\text{Mn}^{2+}$  may be present since  $\text{Mn}^{2+}$  does not form a precipitate with either  $\text{NaCl}$  or  $\text{Na}_2\text{SO}_4$ . A precipitate formed with  $\text{NaOH}$ ; the solution must contain  $\text{Mn}^{2+}$  because it forms a precipitate with  $\text{OH}^-$  [ $\text{Mn}(\text{OH})_2(\text{s})$ ].

4.47



$$0.0750 \text{ L} \times 0.100 \text{ mol AgNO}_3/1 \text{ L} \times 1 \text{ mol Na}_2\text{CrO}_4/2 \text{ mol AgNO}_3 \times$$

$$161.98 \text{ g Na}_2\text{CrO}_4/1 \text{ mol Na}_2\text{CrO}_4 = 0.607 \text{ g Na}_2\text{CrO}_4$$

# Zumdahl Chemistry Chapter 4 Answers

**The Princeton Review**



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