

▶ 실전 연습 해답

제2장

Self-Check Exercise 2.1

$$357 = 3.57 \times 10^2$$

$$0.0055 = 5.5 \times 10^{-4}$$

Self-Check Exercise 2.2

- Three significant figures. The leading zeros (to the left of the 1) do not count, but the trailing zeros do.
- Five significant figures. The one captive zero and the two trailing zeros all count.
- This is an exact number obtained by counting the cars. It has an unlimited number of significant figures.

Self-Check Exercise 2.3

- $12.6 \times 0.53 = 6.678 = 6.7$
 $\begin{array}{r} 12.6 \\ \times 0.53 \\ \hline 378 \\ 630 \\ \hline 6.678 \end{array}$
- $12.6 \times 0.53 = 6.7$
 $\begin{array}{r} 12.6 \\ \times 0.53 \\ \hline 378 \\ 630 \\ \hline 6.678 \end{array}$
- $\begin{array}{r} 25.36 \\ -4.15 \\ \hline 21.21 \end{array}$ $\begin{array}{r} 21.21 \\ \times 2.317 \\ \hline 9.11408 = 9.114 \end{array}$

Self-Check Exercise 2.4

$$0.750 \text{ qt} \times \frac{1.06 \text{ qt}}{1 \text{ qt}} = 0.795 \text{ qt}$$

Self-Check Exercise 2.5

$$225 \frac{\text{mi}}{\text{h}} \times \frac{1760 \text{ yd}}{1 \text{ mi}} \times \frac{1 \text{ yd}}{1.094 \text{ yd}} \times \frac{1 \text{ km}}{1000 \text{ yd}} = 362 \frac{\text{km}}{\text{h}}$$

Self-Check Exercise 2.6

The best way to solve this problem is to convert 172 K to Celsius degrees. To do this we will use the formula $T_{\text{C}} = T_{\text{K}} - 273$.

In this case

$$T_{\text{C}} = T_{\text{K}} - 273 = 172 - 273 = -101$$

So 172 K = -101 °C, which is a lower temperature than -75 °C. Thus 172 K is colder than -75 °C.

Self-Check Exercise 2.7

The problem is $41 \text{ °C} = ? \text{ °F}$.

Using the formula

$$T_{\text{F}} = 1.80(T_{\text{C}}) + 32$$

we have

$$T_{\text{F}} = 1.80(41) + 32 = 74 + 32 = 106$$

That is, $41 \text{ °C} = 106 \text{ °F}$.

Self-Check Exercise 2.8

This problem can be stated as $239 \text{ °F} = ? \text{ °C}$.

Using the formula

$$T_{\text{C}} = \frac{T_{\text{F}} - 32}{1.80}$$

we have in this case

$$T_{\text{C}} = ? \text{ °C} = \frac{239 - 32}{1.80} = \frac{207}{1.80} = 115$$

That is, $239 \text{ °F} = 115 \text{ °C}$.

Self-Check Exercise 2.9

We obtain the density of the cleaner by dividing its mass by its volume.

$$\text{Density} = \frac{\text{mass}}{\text{volume}} = \frac{28.1 \text{ g}}{35.8 \text{ mL}} = 0.785 \text{ g/mL}$$

This density identifies the liquid as isopropyl alcohol.

제3장

Self-Check Exercise 3.1

Items (a) and (c) are physical properties. When the solid gallium melts, it forms liquid gallium. There is no change in composition. Items (b) and (d) reflect the ability to change composition and are thus chemical properties. Statement (b) means that platinum does not react with oxygen to form some new substance. Statement (d) means that copper does react in the air to form a new substance, which is green.

Self-Check Exercise 3.2

- Milk turns sour because new substances are formed. This is a chemical change.
- Melting the wax is a physical change (a change of state). When the wax burns, new substances are formed. This is a chemical change.

Self-Check Exercise 3.3

- Maple syrup is a homogeneous mixture of sugar and other substances dispersed uniformly in water.
- Helium and oxygen form a homogeneous mixture.
- Oil and vinegar salad dressing is a heterogeneous mixture. (Note the two distinct layers the next time you look at a bottle of dressing.)
- Common salt is a pure substance (sodium chloride), so it always has the same composition. (Note that other substances such as iodine are often added to commercial preparations of table salt, which is mostly sodium chloride. Thus commercial table salt is a homogeneous mixture.)

제4장

Self-Check Exercise 4.1

- P_2O_{10} b. UF_6 c. AsCl_3

Self-Check Exercise 4.2

In the symbol $^{90}_{38}\text{Sr}$, the number 38 is the atomic number, which represents the number of protons in the nucleus of a strontium atom. Because the atom is neutral overall, it must also have 38 electrons. The number 90 (the mass number) represents the number of protons plus the number of neutrons. Thus the number of neutrons is $A - Z = 90 - 38 = 52$.

Self-Check Exercise 4.3

The atom $^{201}_{80}\text{Hg}$ has 80 protons, 80 electrons, and $201 - 80 = 121$ neutrons.

Self-Check Exercise 4.4

The atomic number for phosphorus is 15 and the mass number is $15 + 17 = 32$. Thus the symbol for the atom is $^{32}_{15}\text{P}$.

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