Mag 4 2010

Oxidation-Reduction 2Jorksheet 1

Identify the following in the reaction: 2Fe³⁺ + Cl₂ → 2Fe³⁺ + 2Cl (+ mavks)

a. The oxidizing agent	C12	e. The species gaining electrons	Cla
b. The species being oxidized	Fe ²⁺	f. The species losing electrons	Fezr
c. The reducing agent	Fez+	g. The product of exidation	Fe3+
d. The species being reduced	Cl ₂	h. The product of reduction	CI-

2. Use the Standard Reduction Table to answer the following. (+ www.s)

- a. Will Br, cooldize Augo? Explain. Yes, Briss is higher on the table
- b. Give the symbol of an ion that will oxidize Mnoo but not Cross ZY2+
- c. Which is a stronger reducing agent, Sn2+ or Fe2+? Sn2+
- d. Can you safely put a gold ring in acidified dichromate solution? What about acidified bromate solution? Explain. Aurs is higher at the table as a reducing agend than $Cr_2O_3^{2-} + 14H + as an exidizing agend, so it will not read. Aurs will not read with <math>BrO_3^- + bH^2$ wither for same reason

Four metals A, B, C, & D were tested with separate solutions of A²⁺, B²⁺, C²⁺ & D²⁺. Some of the results are summarized in the following table:

*		Solution			
Hetal.	Alle.	Bit.	C2+	D2+	
Α		reaction	- reaction		
.0.				reaction	
D	Cit reaction				

List the ions in order from the strongest to weakest oxidizing agent. (2 marks) BALL

$$C^{z+} + 2e^- \rightarrow C$$
 $A^{z+} + 2e^- \rightarrow A$
 $B^{z+} + 2e^- \rightarrow B$
 $D^{z+} + 2e^- \rightarrow B$

4. Balance the following half reactions. (s wwis)

a. BrO₃
$$\rightarrow$$
 Br (bask)
BrO₃ $^{-}$ +6H⁴ +6e $^{-}$ \rightarrow Br $^{-}$ +3H₂O $^{-}$ BrO₃ $^{-}$ +3H₂O +6e $^{-}$ \rightarrow 6HF $^{-}$ +6OH $^{-}$
b. HClO₂ \rightarrow ClO₃ (acidic)
H₂O + HClO₂ \rightarrow ClO₃ +3H+ +3e $^{-}$

3 Ga+5H20+9e - → H2 GaO3 + 80H-

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