

BOYLE'S LAW

Name _____

Boyle's Law states that the volume of a gas varies inversely with its pressure if temperature is held constant. (If one goes up, the other goes down.) We use the formula:

$$P_1 \times V_1 = P_2 \times V_2$$

Solve the following problems (assuming constant temperature).

1. A sample of oxygen gas occupies a volume of ~~250 mL~~ ^{25L} at ~~740 torr~~ ^{.97 atm} pressure. What volume will it occupy at ~~900 torr~~ ^{1.05 atm} pressure?

2. A sample of carbon dioxide occupies a volume of 3.50 liters at ~~125 kPa~~ ^{1.23 atm} pressure. What pressure would the gas exert if the volume was decreased to 2.00 liters?

3. A 2.0 liter container of nitrogen had a pressure of 3.2 atm. What volume would be necessary to decrease the pressure to 1.0 atm?

4. Ammonia gas occupies a volume of ~~400 mL~~ ^{45L} at a pressure of ~~20 mm Hg~~ ^{2.1 atm}. What volume will it occupy at standard pressure?

5. A ~~125 mL~~ ^{17L} sample of neon had its pressure changed from ~~75 kPa~~ ^{.75 atm} to ~~150 kPa~~ ^{1.5 atm}. What is its new volume?

6. A sample of hydrogen at 1.5 atm had its pressure decreased to 0.50 atm producing a new volume of ~~750 mL~~ ^{45L}. What was its original volume?

7. Chlorine gas occupies a volume of 1.2 liters at ~~720 torr~~ ^{.95 atm} pressure. What volume will it occupy at 1 atm pressure?

8. Fluorine gas exerts a pressure of ~~900 torr~~ ^{1.2 atm}. When the pressure is changed to 1.50 atm, its volume is ~~250 mL~~ ^{25 L}. What was the original volume?

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LL Leslie



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