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# Sullivan Realty Mini Practice Set Answers

**LP Steffe**



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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable  
measure an appropriate volume of each liquid eg 25 cm<sup>3</sup> place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less  
exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of  
sodium hydroxide solution with hcl as 57.9 kJ mol<sup>-1</sup>

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produce salts neutralisation is the reaction between an acid and a base

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic  
eg 1 hcl aq + naoh aq → nacl aq + h<sub>2</sub>o l for which ΔH = 57 kJ mol<sup>-1</sup> eg 2 h<sub>2</sub>no<sub>3</sub> aq + 2koh aq → 2h<sub>2</sub>o l + 2kno<sub>3</sub> aq for which ΔH = 57 kJ mol<sup>-1</sup>

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example  $\text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{o aq}$  you can see that one mole of water is produced when one mole of hcl or one mole of naoh reacts

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