

PAPER 6 – Combined science

Suggest the average length of the three beams is used instead of measuring the length of just one beam. Can identify anomalies (not all beams the same).

Suggest a different measurement the student could take for comparing the size of beams. Mass or volume.

When stating a conclusion for a test mention all the experiments, for example, Beam A contains starch AND Beam B. Don't just mention Beam A and stop.

Suggest how the experiment could be changed so that the highest temperatures reached can be compared more reliably. Repeat the experiment twice for accuracy, use the same initial temperature for all the experiments, along with same mass or amount/ size of the metal.

To find a suitable scale to use on a graph, take both the maximum and minimum value and subtract them, put a bracket, then divide them by the number of large squares you have on your grid. Afterwards, take your answer and multiply it by 2, then you have 2 values, between them is a range and you have to pick the number between them that is divisible by 10. For example, the result 2.1 and 4.2, a suitable scale between them is 2.5.

Both hydrogen carbonate and limewater are indicators that are specific for identifying carbon dioxide.

Describe one technique for measuring a person's rate. Count pulse in 60 seconds, or use a pulse meter/ smart watch.

Make the mass of test tube D. Used as a controlled variable, or to make sure that the city responsible for the results.

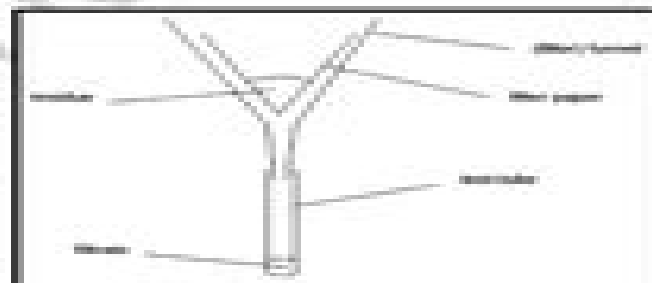
Explain why the test tube was placed in the water bath. To make the temperature constant.

Explain the purpose of using water instead of petroleum in beaker B. As a control, so compare to make petroleum is responsible for the results.

The reaction finishes when the reactants are used up, which is why the rate of the reaction is indicated as a constant value for the remaining of the time.

Assuming that your measuring cylinder is dry when taking volume measurements, does knowing the accuracy of your measurements help.

Suggest why it is better to use the change in temperature to compare the heat loss rather than the final temperature. The starting temperatures are not the same.



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T. A. Glendinning



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