

Multiple Choice: Acid/Base Titrations

1. A solution of known concentration is the definition of a

- A) buffer solution. B) neutral solution. C) standard solution. D) saturated solution

ANS: C

1. A weak acid ($K_a = 4.5 \times 10^{-5}$) is neutralized with a weak base ($K_b = 5.6 \times 10^{-6}$). What is the pH of the resulting solution?

- A) acidic B) basic C) neutral D) pH = 3.2 E) pH = 7

ANS: A

1. The equivalence point of a neutralization reaction is 8.5. What is the best indicator to use?

- A) crystal violet B) bromocresol green C) thymol blue D) methyl red E) alizarin

ANS: C

1. What is the pH of an aqueous solution that is 0.10 M HCOOH ($K_a = 1.8 \times 10^{-4}$) and 0.10 M NaHCO_2 ?

- A) 3.74 B) 10.26 C) 5.62 D) 2.38 E) 5.74

ANS: A

1. Which of the following indicators would be most suitable for the titration of 0.10 M HBr(aq) with 0.10 M KOH(aq) ?

- A) bromothymol blue ($\text{p}K_{in} = 7.1$) B) thymol blue ($\text{p}K_{in} = 1.7$)
C) bromophenol blue ($\text{p}K_{in} = 3.9$) D) alizarin yellow ($\text{p}K_{in} = 11.2$)
E) methyl orange ($\text{p}K_{in} = 3.4$)

ANS: A

3. 200.0 mL of 0.200 M HCl is titrated with 0.050 M NaOH . What is the pH after the addition of 100. mL of the NaOH solution?

- A) 0.93 B) 1.45 C) 1.03 D) 0.76 E) 0.82

ANS: A

1. If a 17.0 mL sample of 0.20 M sulfuric acid (H_2SO_4) is titrated with 0.14 M NaOH , it will require how many mL of titrant to reach neutralization?

- A) 9.7 mL B) 48.6 mL C) 389 mL D) 34.0 mL E) 28.6 mL

ANS: B

1. When NaOH(aq) is titrated with HCl(aq) , what is the pH at the equivalence point?

- A. equal to 7 B. greater than 7 C. less than 7 D. undefined

Titration Multiple Choice

Deepak Chowrasia, Nisha Sharma

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