

## CP6 Radioactivity

### Background Radiation

#### What is natural background radiation?

#### State 2 sources of natural background radiation:

#### Draw a labelled diagram of Rutherford's model State 2 sources of background radiation that does not occur naturally

#### How can radiation be measured and describe how the equipment works.

#### Describe the experiment Rutherford carried out which allowed him to come up with his model.

### Types of radiation

#### Describe the structure of an alpha particle What is the radius of a nucleus?

#### What is the radius of an atom?

#### What is the charge on an alpha particle?

#### How is Rutherford's model of the atom similar to Bohr's model? Describe the structure of a beta particle

#### What is the charge on a beta particle?

### Inside Atoms

#### Draw a labelled diagram of the modern day structure of an atom Describe the structure of a gamma ray

#### What does unstable mean?

	Mass	Charge
Proton		
Neutron		
Electron	Alpha	Beta
Gamma		Weak

Complete the table

Penetration	C	N	Mg	Co
Mass number				
Atomic number				
Protons				
Neutrons				

Explain why alpha particles are more ionising than beta particles

#### Define an isotope

### Radioactive decay

#### Complete the table

Particle	Symbol
Alpha	
Beta	$\beta^-$
Positron	
Neutron	n

#### Describe how a magnesium atom becomes a sodium atom. What happens to the atomic and mass number when a neutron is emitted?

### Electrons and Orbits

#### Radioactive decay happens when an atom absorbs electromagnetic radiation Describe what happens during $\beta^-$ decay and the effect on the nucleus

#### Describe what happens when an atom emits electromagnetic radiation Describe what happens during $\beta^-$ decay and the effect on the nucleus

#### What is the difference between emission and absorption spectra Complete the table summarising what happens to the mass and atomic number during radioactive decay

	Mass number	Atomic number
Alpha		
Beta		Increases by 1

#### Radium has a mass number of 226 and an atomic number of 88.

#### Write a nuclear equation when this nucleus emits an alpha particle

#### Write a nuclear equation if Iodine-131 undergoes $\beta^-$ decay. Iodine has an atomic number of 53 and loses an electron to form an ion. How many protons and electrons are there? Write a nuclear equation if Potassium-37 undergoes $\beta^-$ decay.

# Zgce Inter Science Paper1 Revision

**Jicheng Xie**



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