

Modeled after AP 2013 #21:

1. $2\text{C}_2\text{H}_6(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of reactants consumed when 2.00 mole of $\text{CO}_2(\text{g})$ is produced?
 A. 0.50 mol
 B. 1.50 mol
 C. **4.50 mol**
 D. 5.00 mol
2. $\text{C}_2\text{H}_6(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of reactants consumed when 2.00 mole of $\text{CO}_2(\text{g})$ is produced?
 A. 0.50 mol
 B. 1.00 mol
 C. **3.50 mol**
 D. 7.00 mol
3. $2\text{C}_2\text{H}_6(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of products made when 3.50 mole of $\text{O}_2(\text{g})$ is consumed?
 A. 2.00 mol
 B. 3.00 mol
 C. **5.00 mol**
 D. 10.0 mol
4. $\text{C}_2\text{H}_6(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of products made when 3.50 mole of $\text{O}_2(\text{g})$ is consumed?
 A. 1.00 mol
 B. **2.00 mol**
 C. 4.00 mol
 D. 8.00 mol

Modeled after AP 2014g #48:

3. $4\text{NH}_3(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
A 7.0 mol sample of $\text{NH}_3(\text{g})$ and a 10. mol sample of $\text{O}_2(\text{g})$ are placed in a previously evacuated container and react according to the equation above. How many moles of $\text{NO}(\text{g})$ are produced when one of the reactants has been totally consumed?
 A. 2.5 mol
 B. **7.0 mol**
 C. 8.0 mol
 D. 28 mol
6. $\text{C}_2\text{H}_6(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{CH}_4(\text{g})$
Samples containing 4.0 mol $\text{C}_2\text{H}_6(\text{g})$ and 5.0 mol $\text{H}_2(\text{g})$ are placed in a previously evacuated container and allowed to react according to the equation above. When one of the reactants is completely consumed, how much product is formed?

Zumdahl Ap Chemistry Review Questions Answers

Mina Lebitz



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