



Atomic Mass of Cadmium (Cd)

Objectives: To isolate the "isotopes" of Cadmium and calculate its atomic mass.

Materials: - Sample of Cadmium (in a baggie)

- a balance

- a calculator

Introduction:

As we now know, Dalton's idea that all atoms of one element were identical is actually incorrect. A sample of any element is a mixture of its various isotopes. Some concepts we will be investigating include how the **average atomic mass** of an element relates to the **atomic mass** and **relative abundance** of its isotopes. As each element is comprised of different isotopes, your sample of cadmium is comprised of different cadmium with different masses. Each type of cadmy represents a different isotope of cadmium.

Pre-Lab Questions:

1. What are isotopes?
2. Define average atomic mass and relative abundance in two complete sentences.

Procedure:

1. Obtain a sample of cadmium that contains three different brands of candy. Treat each brand of candy as an isotope of cadmium.
2. On a piece of paper towel, separate the three isotopes into groups labeled A, B and C, and measure the total mass of each isotope. Record the masses of each isotope in the **Total Mass** row of the data table.
3. Count the number of atoms in each sample. Record the number in the **numbers** row.
4. Follow the calculations described below to understand how to fill out the remainder of the data table.

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James F. Hall



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