

8.2**THE NATURE OF COVALENT BONDING****Section Review****Objectives**

- State a rule that usually tells how many electrons are shared to form a covalent bond.
- Describe how electrons are formulated.
- Predict when two atoms are likely to be joined by a double or a triple covalent bond.
- Distinguish between a single covalent bond and other covalent bonds.
- Describe how the strength of a covalent bond is related to its bond dissociation energy.
- Describe how resonance structures explain bonding.
- Identify some exceptions to the octet rule.

Vocabulary

- | | | |
|------------------------|----------------------------|----------------------------|
| • single covalent bond | • double covalent bonds | • polyatomic ion |
| • structural formulas | • triple covalent bonds | • bond dissociation energy |
| • unshared pairs | • coordinate covalent bond | • resonance structures |

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are presented in this section. Each blank can be completed with a term, short phrase, or number.

- When atoms share electrons to gain the ____1____ configuration of a noble gas, the bonds formed are ____2____. A ____3____ pair of valence electrons constitutes a ____4____ covalent bond. Pairs of valence electrons that are not shared between atoms are called ____5____. Sometimes two or three pairs of electrons may be shared to give ____6____ covalent bonds. In some cases, only one of the atoms in a bond provides the pair of bonding electrons; this is a ____7____ ____8____. It is required to break covalent bonds between atoms. The total energy required to break the bond between two covalently bonded atoms is known as the ____9____.
- When it is possible to write two or more valid electron dot formulas for a molecule or ion, each formula is referred to as a ____10____.

The Nature Of Covalent Bonding Answer Key

K Morrison



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Table of Contents The Nature Of Covalent Bonding Answer Key

1. Understanding the eBook The Nature Of Covalent Bonding Answer Key
 - The Rise of Digital Reading The Nature Of Covalent Bonding Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying The Nature Of Covalent Bonding Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an The Nature Of Covalent Bonding Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from The Nature Of Covalent Bonding Answer Key
 - Personalized Recommendations
 - The Nature Of Covalent Bonding Answer Key User Reviews and Ratings
 - The Nature Of Covalent Bonding Answer Key and Bestseller Lists
5. Accessing The Nature Of Covalent Bonding Answer Key Free and Paid eBooks

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 - The Nature Of Covalent Bonding Answer Key eBook Subscription Services
 - The Nature Of Covalent Bonding Answer Key Budget-Friendly Options
6. Navigating The Nature Of Covalent Bonding Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - The Nature Of Covalent Bonding Answer Key Compatibility with Devices
 - The Nature Of Covalent Bonding Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of The Nature Of Covalent Bonding Answer Key
 - Highlighting and Note-Taking The Nature Of Covalent Bonding Answer Key
 - Interactive Elements The Nature Of Covalent Bonding Answer Key
8. Staying Engaged with The Nature Of Covalent Bonding Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers The Nature Of Covalent Bonding Answer Key
9. Balancing eBooks and Physical Books The Nature Of Covalent Bonding Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection The Nature Of Covalent Bonding Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine The Nature Of Covalent Bonding Answer Key
- Setting Reading Goals The Nature Of Covalent Bonding Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of The Nature Of Covalent Bonding Answer Key
- Fact-Checking eBook Content of The Nature Of Covalent Bonding Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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